

- p_s = pressure of saturated liquid.
 psig = pressure pounds per square inch, gage.
 psia = pressure pounds per square inch, absolute.
 Q = quantity of heat, Btu.
 Q_c = heat loss from condenser, Btu per pound refrigerant.
 Q_i = heat dissipated in cooling water, Btu per hour.
 s = entropy.
 Δs = entropy change between suction and discharge.
 T = absolute temperature, Fahrenheit degrees.
 T_{avg} = average temperature, Fahrenheit degrees, absolute, of gas passing through compressor.
 T_c = condenser temperature, Fahrenheit degrees, absolute.
 T_e = evaporator temperature, Fahrenheit degrees, absolute.
 (TVE) = total volumetric efficiency.
 t_{sd} = degrees superheat at discharge condition of vapor leaving compressor.
 t_d = discharge temperature, Fahrenheit degrees.
 V_o = clearance, percentage of volume, swept by piston, which is contained in spaces at end of cylinder when piston is at end of stroke (clearance includes valve spaces, etc.).
 v_s = specific volume of gas at suction, cubic feet per pound.
 v_d = specific volume of gas at discharge, cubic feet per pound.
 W_r = refrigerant rate, pounds per minute.
 x = proportion of liquid in mixture of vapor and liquid.

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CHAPTER 37

DEHUMIDIFICATION BY SORBENT MATERIALS

Definitions and Principles, Adsorbents, Dehumidification by Solid Adsorbents, Dehumidification Equipment Using Solid Adsorbents, Absorbents, Dehumidification by Liquid Absorbents, Dehumidification Equipment Using Liquid Absorbents, Calculation of Moisture Load, Vapor Transfer to Dehumidified Space

DEHUMIDIFICATION as used herein is the reduction of the water vapor content of a given volume of air or other gas. The term thus describes a special case of dehydration which covers the removal of moisture in any form from matter. The degree of dehumidification required varies greatly with different applications, and is one of the prime considerations influencing the choice of a method. Dehumidification may be accomplished by latent heat removal, together with sensible heat removal, as described in Chapters 29, 34, and 35, or by the use of sorbents.

Sorbents are substances which have the property of extracting and holding other substances (usually gases or vapors, e.g., water vapor), brought into contact with them. All materials are sorbents to a greater or lesser degree. The weight of water held by a substance will increase or decrease, depending upon whether the vapor pressure of the water held by the substance is less or greater, respectively, than the partial pressure of water vapor in the surrounding atmosphere. As generally used, however, the term sorbents refers to those materials having a capacity for moisture which is large compared to their volume and weight. Such materials are divided into two general classifications:

1. *Adsorbent*—A sorbent which does not change physically or chemically during the sorption process. Certain solid materials, such as activated alumina, silica gel, activated bauxites, and activated charcoal have this property. The action of adsorbents, most of which adsorb some gases and condensable vapors besides water vapor, is selective. Thus, in the case of a mixture containing both water and organic vapors, silica gel would remove the water vapor in preference to the organic vapors, while the reverse would be true in the case of activated carbon. The selective property of adsorbents is made use of in some instances for the removal of objectionable and contaminating vapors from an air or gas mixture. (See Chapter 8.)
2. *Absorbent*—A sorbent which changes either physically, chemically, or both, during the sorption process. Calcium chloride is an example of a solid absorbent, while liquid absorbents include solutions of lithium chloride, calcium chloride, lithium bromide, and the ethylene glycols.

ADSORBENTS

The ability of an adsorbent to remove water vapor from a gas is explained by the fact that the vapor pressure of the water in the adsorbent (when in the reactivated condition) is less than the partial pressure of the water vapor in the surrounding atmosphere. For instance, when an active adsorbent is brought into contact with a gas of high humidity, there is a tendency for the vapor pressure of the water in the adsorbent to reach equilibrium with the partial pressure of the water in the surrounding gas, with