

of organic substances such as tobacco, wood, coal, oil, tar and other carbonaceous materials. The term *smoke* is commonly applied to the mixture of solid, liquid and gaseous products of combustion, although the technical literature prefers to distinguish between such components as soot or carbon particles, fly-ash, cinders, tarry matter, unburned gases, and gaseous combustion products. The finest particulate constituents are much less than 1 micron in size, often in the range of 0.1 to 0.3 micron.

Mists and Fogs are liquid particulate air contaminants.

Mists are very small airborne droplets of materials that are ordinarily liquid at normal temperatures and pressures. They may be formed by atomizing, spraying, splashing, mixing, violent chemical reaction, electrolytic evolution of gas from a liquid, or escape of a dissolved gas upon release of pressure. Very small droplets expelled or atomized into the air by sneezing constitute mists containing microorganisms that become air contaminants.

Fogs are limited by some classifications to airborne droplets formed by condensation from the vapor state. This arbitrary distinction between mist and fog is of minor importance, as both terms are used to indicate the particulate state of airborne liquids (occasionally termed *aerosols*). *Fog nozzles* are so named because of their ability to produce extra fine droplets as compared to the mist from ordinary spray devices. The highly volatile nature of some liquids quickly reduces their airborne droplets from the mist to the fog range, and eventually to the vapor phase until the air becomes saturated with that liquid. Many droplets in fogs or clouds are microscopic and submicroscopic in size, and may be conceived as the transition state between the larger mists and the vapors.

Vapors and Gases are non-particulate air contaminants.

Vapors are the gaseous phase of substances that are either liquid or solid in their commonly known state, examples being gasoline, kerosene, benzene, carbon tetrachloride, mercury, iodine, camphor. Vapors may be changed to the solid or liquid form by increasing the pressure, decreasing the temperature or applying both processes simultaneously. They are removed from the air by condensation with less difficulty than are the gases.

Gases are normally formless fluids which tend to occupy a space or enclosure completely and uniformly at ordinary temperatures and pressures. The following substances qualify as gases: oxygen, nitrogen, carbon dioxide, carbon monoxide, hydrogen, ammonia, sulfur dioxide. Gases, likewise, may be solidified or liquefied by the proper control of temperature and pressure.

The preceding classification is not suitable for the airborne living organisms, which range in size from the submicroscopic viruses to the largest pollen grains, not considering the smallest insect life. Bacteria range from about 0.2 to 5 microns in size, fungus spores from 1 to 10 microns, and pollen from 5 to 150 microns.

SIZES OF AIRBORNE PARTICLES

Fig. 1 is a graphic tabulation of the properties of airborne solids and liquids arranged according to size on the micron scale. There are 25,400 microns in 1 inch.

Particles larger than 10 microns are unlikely to remain suspended in air currents of moderate strength, but settle out by gravity at speeds dependent upon the shape, size and specific gravity of the particle, wind velocity, orientation of the collecting surface, and topography. These larger particles are of major interest to the engineer in the solution of nuisance problems, but it is usually the smaller particles, or those below 10 microns, that remain in the air long enough to be of hygienic as well as economic significance.

Industrial dust particles are predominantly of the order of 1 micron in size. Tremendous numbers are also present in the submicroscopic range below 0.5 micron, but those below 0.1 micron are not believed at

present to be of practical importance, possibly due to their exceedingly small mass in comparison with the balance of airborne matter. In fact, particles this small may become the *permanent* atmospheric impurities that have little, if any, opportunity of settling because of the continual motion imparted to them by air currents and the molecular activity of gases (Brownian Movement).

DIAM. OF PARTICLES IN MICRONS	SCALE OF ATMOSPHERIC IMPURITIES	RATE OF NUMBER SURFACE SETTLING OF PAR- AREA IN IN F.R.M. TILES IN SQUARE FOR ONE CUBIC INCHES SPHERES AIR CONTAINING OF 1,000 GRAMS OF IMPURITIES PER CUBIC FT. (DENSITY = 1) AT 70° F.		LAWS OF SETTLING IN RELATION TO PARTICLE SIZE (LINES OF DEMARCATION APPROX.)	
		PARTICLES FALL WITH INCREASING VELOCITY		PARTICLES SETTLE WITH CONSTANT VELOCITY	
8000	1" RAIN	1750		$C = \sqrt{\frac{2gd_s}{3Ks_z}}$	C-Velocity cm/sec
6000	1" DRIZZLE	790	.075	$C = 24.9\sqrt{ds}$	C-Velocity ft/min.
4000	1" FOG	555	.6	$C = \frac{2r^2g}{9}$	d-Diam. of particle in cm.
2000	1" MIST	555	.6	$C = \frac{2r^2g}{9}$	D-Diam. of particle in microns
1000	1" MIST	555	.6	$C = \frac{2r^2g}{9}$	r-Radius of particle in cm.
800	1" MIST	555	.6	$C = \frac{2r^2g}{9}$	g=981 cm/sec. ²
600	1" MIST	555	.6	$C = \frac{2r^2g}{9}$	s _p =Density of particle
400	1" MIST	555	.6	$C = \frac{2r^2g}{9}$	s _a =Density of Air
200	1" MIST	555	.6	$C = \frac{2r^2g}{9}$	C=300,460 s _a d ³
100	1" MIST	555	.6	$C = \frac{2r^2g}{9}$	C=.00592 s _a d ³
80	1" MIST	555	.6	$C = \frac{2r^2g}{9}$	η=Viscosity of air in poises
60	1" MIST	555	.6	$C = \frac{2r^2g}{9}$	λ=10 ⁻⁵ cm.
40	1" MIST	555	.6	$C = \frac{2r^2g}{9}$	(Mean free path of gas molecules)
20	1" MIST	555	.6	$C = \frac{2r^2g}{9}$	λ=10 ⁻⁵ cm.
10	1" MIST	555	.6	$C = \frac{2r^2g}{9}$	(Mean free path of gas molecules)
8	1" MIST	555	.6	$C = \frac{2r^2g}{9}$	(Mean free path of gas molecules)
6	1" MIST	555	.6	$C = \frac{2r^2g}{9}$	(Mean free path of gas molecules)
4	1" MIST	555	.6	$C = \frac{2r^2g}{9}$	(Mean free path of gas molecules)
2	1" MIST	555	.6	$C = \frac{2r^2g}{9}$	(Mean free path of gas molecules)
1	1" MIST	555	.6	$C = \frac{2r^2g}{9}$	(Mean free path of gas molecules)
.8	1" MIST	555	.6	$C = \frac{2r^2g}{9}$	(Mean free path of gas molecules)
.6	1" MIST	555	.6	$C = \frac{2r^2g}{9}$	(Mean free path of gas molecules)
.4	1" MIST	555	.6	$C = \frac{2r^2g}{9}$	(Mean free path of gas molecules)
.2	1" MIST	555	.6	$C = \frac{2r^2g}{9}$	(Mean free path of gas molecules)
.1	1" MIST	555	.6	$C = \frac{2r^2g}{9}$	(Mean free path of gas molecules)
.01	1" MIST	555	.6	$C = \frac{2r^2g}{9}$	(Mean free path of gas molecules)
.001	1" MIST	555	.6	$C = \frac{2r^2g}{9}$	(Mean free path of gas molecules)

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FIG. 1. SIZES AND CHARACTERISTICS OF AIRBORNE PARTICULATE MATTER

The survey¹ of atmospheric pollution in 14 American cities conducted from 1931 to 1933 indicated the average size of outdoor dust particles to be 0.5 micron, as collected by the Owens jet dust counter and measured under the microscope. Inability of the *light field* microscope to reveal