

I have been given a review of Standard 1910.93a Vinyl Chloride.

Signed: Billy R Patterson

Date: 9/17/76

Review Next: 9-17-76

PERSONNEL WORK SHEET

DATE: 6-2-76

NAME: Billy Patterson

PUMP: 1056

TUBE: #8

PUMP FACTOR: 0.62

TYPE OF SAMPLE: 15 min

JOB ANALYSIS: Breaking down
and sampling a UCL
Tank car on C.T.R.

A = Chromatograph Factor: .000107

B = Peak Ht. of Sample: 0

C = Vol. of CS₂ Used: 20 mls.

D = Efficiency: 100%

E = Vol. of Air Pumped: 3.8

F = M.W. of Compound: 62.5

Constants = 22.4 L/Mole and 1000

Pump Stop # 377029
Pump Start # - 370894
PUMP UNITS = 6135

Pump Units 6135 Liters 3803 cc = 3.8
Pump Factor X 0.62
C.C. OF AIR = 3803
1000

FORMULA: ppm (X) = $\frac{(A)(B)(C)(22.4\text{L/Mole})(1000)}{(M.W)(D)(E)}$

ppm = $\frac{(A)(0)(448,000)}{(62.5)(1)(3.8)}$

ppm = 0

ppm = 0 or N.D.

PERSONNEL WORK SHEET

DATE: 6-2-76

NAME: Billy Patterson

PUMP: 1055

TUBE: # 13

PUMP FACTOR: 0.580

TYPE OF SAMPLE: 15 MIN

JOB ANALYSIS: Breaking down
and sampling a UCL
Tank car on C.C.R.

A = Chromatograph Factor: 1,000/107

B = Peak Ht. of Sample: 1.5 DIV

C = Vol. of CS₂ Used: 20 mls.

D = Efficiency: 100 %

E = Vol. of Air Pumped: 0.63 Liters

F = M.W. of Compound: 62.5

Constants = 22.4 L/Mole and 1000

Pump Stop # 100052
Pump Start # - 098973
PUMP UNITS = 1079

Pump Units 1079
Pump Factor X 0.58
C.C. OF AIR = 625.8

Liters 625.8 cc = 0.63
1000

FORMULA: ppm (X) = $\frac{(A)(B)(C)(22.4\text{L/Mole})(1000)}{(M.W)(D)(E)}$

ppm = $\frac{(1,000/107)(1.5)(448,000)}{(62.5)(1)(0.63)}$

ppm = 72.1
39.3

ppm = 1.8