

TABLE 7. GENERAL DATA OF COMBUSTIBLE ELEMENTS AND COMPOUNDS*

SUBSTANCE	MOLECULAR SYMBOL	CHEMICAL REACTION OF COMBUSTION	IGNITION TEMPERATURE, °F. DEG.	CALORIFIC VALUE		THEORETICAL OXYGEN AND AIR REQUIREMENTS			
				Btu per Lb		Lb per Lb		Cu Ft per Cu Ft	
				Gross	Net	O ₂	Air	O ₂	Air
Carbon (to CO)	—	2C + O ₂ = 2CO	—	3950	—	1.33	5.76	—	—
Carbon (to CO ₂)	—	2C + 2O ₂ = 2CO ₂	—	14093	—	2.66	11.53	—	—
Sulfur (to SO ₂)	—	S + O ₂ = SO ₂	—	3983	—	1.00	4.29	—	—
Sulfur (to SO ₃)	—	2S + 3O ₂ = 2SO ₃	—	5940	—	1.50	6.43	—	—
Carbon Monoxide	CO	2CO + O ₂ = 2CO ₂	1166-1319	4347	—	0.57	2.47	0.5	2.39
Methane	CH ₄	CH ₄ + 2O ₂ = CO ₂ + 2H ₂ O	1260-1380	23879	21520	3.99	17.27	2.0	9.53
Acetylene	C ₂ H ₂	2C ₂ H ₂ + 5O ₂ = 4CO ₂ + 2H ₂ O	763-824	21500	20776	3.07	13.30	2.5	11.91
Ethylene	C ₂ H ₄	C ₂ H ₄ + 3O ₂ = 2CO ₂ + 2H ₂ O	986-1123	21644	20295	3.42	14.81	3.0	14.29
Ethane	C ₂ H ₆	2C ₂ H ₆ + 7O ₂ = 4CO ₂ + 6H ₂ O	990-1120	22320	20432	3.73	16.12	3.5	16.68
Hydrogen	H ₂	2H ₂ + O ₂ = 2H ₂ O	1063-1166	60958 ^c	51571 ^c	7.94	34.34	0.5	2.39
Hydrogen Sulfide	H ₂ S	2H ₂ S + 3O ₂ = 2H ₂ O + 2SO ₂	599-608	7100	6545	1.41	6.10	1.5	7.15
Propane	C ₃ H ₈	C ₃ H ₈ + 5O ₂ = 3CO ₂ + 4H ₂ O	950-1080	21661	19944	3.63	15.70	5.0	23.82
n-Butane	C ₄ H ₁₀	2C ₄ H ₁₀ + 13O ₂ = 8CO ₂ + 10H ₂ O	890-1020	21308	19680	3.58	15.49	6.5	30.97
Commercial Propane	—	—	920-1020	21560	19865	3.60	15.58	4.9	23.4
Commercial Butane	—	—	900-1000	21180	19591	3.54	15.3	6.3	30.0

* Values in table taken chiefly from page 61 of Fuel Plus Gases published by American Gas Association.

b Gas measured at 80 F and 30 in. Hg.

c Value from National Bureau of Standards.

carbon may unite with oxygen to form carbon monoxide, CO, and some of the hydrogen and hydrocarbon gases may not be burned at all. When carbon monoxide or other combustible gases are present in the flue gases, there is a loss of heat produced per unit of fuel consumed, and a lower combustion efficiency is obtained. Incomplete combustion may result from any or all of the following three conditions: (1) inadequate air supply, (2) insufficient mixing of air and gases, and (3) a temperature too low to produce ignition or maintain combustion.

AIR REQUIRED FOR PERFECT COMBUSTION

Air requirements for combustion of solid and liquid fuels are ordinarily expressed in pounds. On the other hand similar requirements for gaseous fuels are usually stated on a cubic foot basis. For solid and gaseous fuels this method of treatment corresponds to the standards of measurement commonly employed by these two industries.

The weight of air required for perfect combustion per pound of solid or liquid fuel may be calculated after substitution of the proper percentages by weight of the various elements obtained from an ultimate analysis of the fuel in Equation 2. For gaseous fuels see Equation 3 for volume of air required.

Solid or Liquid Fuels:

$$\text{Pounds air required per pound fuel} = \frac{34.56}{100} \left[\frac{C}{3} + \left(H - \frac{O}{8} \right) + \frac{S}{8} \right] \quad (2)$$

For Gaseous Fuels:

$$\begin{aligned} \text{Cubic feet air required per cubic foot gas} = & 2.39 (CO + H_2) + 9.53 CH_4 + \\ & 16.68 C_2H_6 + 23.82 C_3H_8 + 30.97 C_4H_{10} + 11.91 C_2H_2 + 14.29 C_2H_4 + \\ & 7.15 H_2S - 4.78 O_2 + 30.47 \text{ Illuminants} \end{aligned} \quad (3)$$

Gaseous fuels may contain a wide variety of components classified as *illuminants*, which are not separated by the usual methods of gas analysis. The principal ones in addition to ethylene and acetylene which are included in Equation 3, with the air required per cubic foot of gas are: propylene, 21.44; butylene, 28.58; pentene, 35.73; benzene, 35.73; toluene, 42.88; and xylene, 50.02. Since toluene and xylene are normally scrubbed from the gas before distribution, they may be disregarded in computing air required for a fuel gas. An approximate value of 30.47 (as shown in Equation 3) may, therefore, be employed. If ethylene and acetylene are included as *illuminants*, it is suggested that the value 19.65 be used.

If it is desired to make the gas calculations on a weight basis the equation is expressed as follows:

$$\begin{aligned} \text{Pounds air required per pound fuel} = & 2.47 CO + 34.34 H_2 + 17.27 CH_4 \\ & + 16.12 C_2H_6 + 15.70 C_3H_8 + 15.49 C_4H_{10} + 13.30 C_2H_2 + \\ & 14.81 C_2H_4 + 6.10 H_2S - 4.32 O_2 \end{aligned} \quad (4)$$

Where approximate results only are desired, values appearing in Table 8 may be substituted for Equations 2, 3, and 4; or the air required for perfect combustion may be estimated by assuming that 0.9 cu ft air is required per 100 Btu of fuel.

Where extreme precision is involved it is suggested that the reader refer to scientific literature published on the subject of combustion and related