

in accordance with the regulations of the National Fire Protection Association, and other sources of ignition, such as welding, must be controlled

Picric acid was once used extensively as a dressing for burns. This use has been largely discontinued because many cases of skin irritation and systemic poisoning resulted from application of concentrations as low as 1 per cent to large areas of the skin.

Picric acid is rapidly absorbed through the unbroken skin and even more rapidly through wounds, leading to headache, shivering, fever, and insomnia. Absorbed in larger quantities, it produces jaundice and severe inflammation of the kidneys. Exposure to the dust of picric acid may cause irritation of the nose and throat, and especially of the eyes, leading to ulceration of the cornea.

See Industrial Data Sheet 351

Poison oak, poison ivy, and poison sumac are plant species whose toxic irritant has been identified as urushiol. The dermatitis is the result of hypersensitization to this compound. Individual susceptibility varies greatly and also changes from time to time, depending, at least in part, on how long it has been since the last exposure. No one is so completely immune, however, that a sufficiently rigorous exposure will not result in dermatitis.

Since reaction to these plants is a hypersensitization, it seems logical that individuals could be desensitized by injections of graduated doses of the purified extract. This has been attempted in many instances, and there are a number of extracts on the market for the purpose. They seem to give some limited protection to sensitive individuals.

The poisoning can most effectively be prevented by destruction of the plants with one or two applications of a herbicide such as one of the 2,4-D formulations. Destruction by grubbing out and burning is both expensive and hazardous for the men doing the work and is no more effective than application of the herbicide.

See Industrial Data Sheet 304

Potassium chlorate—see Sodium chlorate.

Potassium cyanide—see Sodium cyanide

Potassium hydroxide is a white solid extremely soluble in water. It is most commonly furnished as a concentrated solution in tank-

car quantities and as the solid in flakes, pellets, and sticks, or cast solid in glass or in metal drums.

The most common injuries suffered with potassium hydroxide are burns of the skin or eyes, caused by solid particles and splashing of the concentrated solution from attempts to dissolve the solid in hot water. The solid alkali generates so much heat on dissolving that there is often local boiling even when cold water is used.

To dissolve the material from a drum in which it is cast solid, the drum should be suspended near the top of a tank of lukewarm water so that the water can circulate through the drum. The high density of the concentrated solution will promote circulation from the top to the bottom of the tank. This method will provide both faster solution and less possibility of injury than any other.

First aid for alkali burns consists in washing away the solid or the solution with large quantities of cold water, preferably under a deluge shower, and treating the injury as a heat burn. The washing of alkali burns should be particularly thorough because of the tendency of the alkalis to penetrate into the skin and continue burning even when thought to be washed away.

Eyes injured by potassium hydroxide should be washed with a gentle stream of water for at least fifteen minutes and should then be immediately brought under medical care. Because of the penetrating properties of alkalis, the injuries may be extremely severe.

Propyl acetate, either a n-propyl or isopropyl, is a mild irritant to the eyes, nose, and upper respiratory tract. These compounds are also potent narcotics, the narcotic effect being about the same as that of ether at the same volume concentration. There is no evidence that the effects are cumulative.

Propyl alcohol (propanol) is mainly a respiratory and eye irritant. Liver damage in animals following repeated inhalations of light concentrations of propyl alcohol has been reported. Isopropyl alcohol is more common in industrial use than n-propyl.

In the body, n-propyl alcohol is oxidized and excreted somewhat more rapidly than is ethyl alcohol, while isopropyl alcohol is excreted considerably more slowly. For this