

71 - 70 20.00 171 ppm A42

43 270 water samples #42 + 107, 108, 109
from which we have been unable to
obtain indications are that the
is in the ppb. range

42 700.00 1000 - A42 present - sample
analysis concentration

44 700.00 water contain 1.43 ppb. A42

SVO Ridge

$$\times 10^{-2} \mu\text{g}$$

$$\frac{1.06}{0.89}$$

$$\times 10^{-2} \mu\text{g}$$

$$\frac{1.10}{1.09} \frac{6.3}{6.2}$$

29N
min (H₂O)

$$\frac{0.56}{0.52}$$

$$\frac{0.60}{0.60} \text{ 1L}$$

mass found

$$0.82 - 1254 \text{ 25g}$$

SVO water interferences

interferences -

$$\frac{1.10 \times 10^{-2} \mu\text{g}}{5 \text{ ml}} = \frac{1.10 \text{ g}}{\text{L}} = \frac{1.10 \times 10^{-5} \mu\text{g}}{\text{ml}} \times 500$$

$$\frac{5.50 \times 10^{-3} \mu\text{g}}{6.3} \times 1000 = \frac{5.5}{6.3} \times 10^{-6} \frac{\mu\text{g}}{\text{kg}}$$

$$\frac{5.70 \times 10^{-3} \mu\text{g}}{6.2 \text{ g}} \times 1000 = \frac{5.7}{6.2} \times 10^{-6} \frac{\mu\text{g}}{\text{kg}}$$

$$\times 10^3 = \frac{5.5}{6.3} \times 10^{-3} \text{ PPB}$$

call
Larry Williams
#4983

Dr. Breck
tomorrow
SAS tom. night.
==

1-3PM
T469

mass water

$$\begin{array}{r} .0285 \\ 20 \overline{) 5700} \\ \underline{40} \\ 170 \end{array}$$

$$\begin{array}{r} 2.85 \\ 1 \overline{) 4.35} \end{array}$$

$$\begin{array}{r} 2.16 \\ 10 \overline{) 216} \end{array}$$

$$2.57 \times 10^{-2} \mu\text{g} / 20 \mu\text{L} = 0.0285 \times 10^{-2} \mu\text{g} / \mu\text{L}$$

$$2.85 \times 10^{-4} \mu\text{g} / \mu\text{L} = 2.85 \times 10^{-4} \frac{\text{g}}{\text{L}} = 2.85 \times 10^{-7} \frac{\text{g}}{\text{mL}}$$

$$\times 5 = 14.25 \times 10^{-7} \frac{\text{g}}{\text{L}} \times 10^3 \frac{\text{mL}}{\text{g}} = 14.25 \times 10^{-4} \frac{\text{mL}}{\text{L}}$$

$$= 14.25 \times 10^{-4} \text{ ppm} \times 10^3 \frac{\text{ppb}}{\text{ppm}} = 14.25 \times 10^{-1} \text{ ppb}$$

$$> = 1.425 \text{ ppb}$$

$$1.08 \times 10^{-2} \mu\text{g} / 5 \mu\text{L} = 0.216 \times 10^{-2} \mu\text{g} / \mu\text{L} = 2.16 \times 10^{-3} \frac{\text{g}}{\text{L}}$$

$$2.16 \times 10^{-6} \frac{\text{g}}{\text{mL}} \times 500 \times 10^2 \text{ mL} = 10.80 \times 10^{-4} \text{ g} / 6.3 \text{ g}$$

$$10.80 \times 10^{-4} \text{ g} / 6.3 \times \frac{1000 \text{ g}}{\text{kg}} = 1.71 \times 10^{-1} \text{ g} / \text{kg}$$

$$\rightarrow \times \frac{10^3 \text{ mg}}{\text{g}} = 171 \text{ ppm}$$

$$\begin{array}{r} 1.71 \\ 63 \overline{) 108000} \\ \underline{63} \\ 450 \\ \underline{420} \\ 30 \\ \underline{21} \\ 90 \\ \underline{63} \\ 270 \end{array}$$