TABLE 1. GENERAL DATA OF COMBUSTIBLE ELEMENTS AND COMPOUNDS

Substance	Molecular Symbol	Chemical Reaction of Combustion	Ignition Temperature[a] Deg F	Calorific Value: Btu per Lb, Higher	Calorific Value: Btu per Lb, Lower	Calorific Value: Btu per Cubic Foot, Higher	Theoretical Oxygen and Air Requirements: Lb per Lb, O_2	Lb per Lb, Air	Cubic Ft per Cubic Ft, O_2	Cubic Ft per Cubic Ft, Air
Carbon (to CO)	—	$2C + O_2 = 2CO$	—	4380	—	—	1.333	5.76	—	—
Carbon (to CO_2)	—	$2C + 2O_2 = 2CO_2$	—	14540	—	—	2.667	11.52	—	—
Sulphur	S_2	—	—	—	—	—	1.000	4.32	—	—
Sulphur (to SO_2)	—	$S + O_2 = SO_2$	—	4050	—	—	—	—	—	—
Sulphur (to SO_3)	—	$2S + 3O_2 = 2SO_3$	—	5940	—	—	—	—	—	—
Carbon monoxide	CO	$2CO + O_2 = 2CO_2$	1166-1319	4380	—	342	0.572	2.46	0.5	2.391
Methane	CH_4	$CH_4 + 2O_2 = CO_2 + 2H_2O$	1202-1346	23850	21670	1073	4.000	17.28	2.0	9.564
Acetylene	C_2H_2	$2C_2H_2 + 5O_2 = 4CO_2 + 2H_2O$	763-824	21460	21020	1590	3.077	13.29	2.5	11.955
Ethylene	C_2H_4	$C_2H_4 + 3O_2 = 2CO_2 + 2H_2O$	986-1123	21450	20420	1675	3.429	14.81	3.0	14.346
Ethane	C_2H_6	$2C_2H_6 + 7O_2 = 4CO_2 + 6H_2O$	986-1123	22230	20500	1883	3.733	16.13	3.5	16.737
Hydrogen	H_2	$2H_2 + O_2 = 2H_2O$	1063-1166	62000	52920	348	8.000	34.56	0.5	2.391
Hydrogen sulphide	H_2S	$2H_2S + 3O_2 = 2H_2O + 2SO_2$	599-608	—	—	—	1.412	6.10	1.5	7.173

[a]From International Critical Tables, 1927.

ignition temperature before it will unite in chemical combination with oxygen to produce combustion. The ignition temperatures for several of the combustible constituents of fuels are presented in Table 1.

HEAT OF COMBUSTION

As previously stated, the process of combustion results in the evolution of heat. The heat generated by the complete combustion of a unit of fuel is constant for a given combination of combustible elements and compounds, and is known as the *heat of combustion*, *calorific value*, or *heating value* of the fuel.

The heat of combustion of the several fuel elements and compounds in their *pure* state is given in Table 1.

The reaction of the carbon in the fuel with oxygen may result in the formation of carbon monoxide or carbon dioxide. In burning to carbon monoxide, the carbon is not completely oxidized and, as shown by the data, the heat produced is considerably less than if it were completely oxidized. This fact is of greatest importance in considering the efficiency of combustion.

The calorific value of a fuel is determined by direct measurement of the heat evolved during combustion in a calorimeter. Although the ash and moisture content of coal from a given mine or locality may vary widely, the heating value of the coal, on a *moisture and ash free* basis, remains relatively constant. It is therefore possible to approximate the heating value of a shipment of coal *as received* if its moisture and ash content are determined, and if the heating value of similar coal on a moisture and ash free basis is known. This may be calculated by Equation 1.

$$\text{Heating value, } \textit{as received} = \frac{\text{Heating value, } \textit{moisture and ash free} \times [100 - (\text{Moisture} + \text{Ash})]}{100} \quad (1)$$

where, moisture and ash are expressed in per cent.

Typical analyses of the coals of the United States are given in *U. S. Bureau of Mines Bulletin* 446.

As practically all fuels contain hydrogen, water vapor is one of the products of combustion. The amount of water vapor produced is proportional to the hydrogen content of the fuel.

When the heating value of a fuel is determined in a calorimeter the water vapor is condensed and the latent heat of vaporization is included in the heating value of the fuel. The heating value so determined is termed the *gross* or *higher* heating value and this is what is ordinarily meant when the heating value of a fuel is specified. In burning the fuel, however, the products of combustion are not cooled to the dew-point and the higher heating value cannot be obtained.

FLAME

The appearance of the flame or products of combustion may serve as an approximate measure of the temperatures developed in the combustion process. The luminosity of a flame is caused by the heating to incandescence of unconsumed particles of combustible matter in the gases, and