

LEAD POLLUTING BY CORROSION

425 LEADERSHIP AVENUE

NEW YORK 17, N. Y.

HEAVY METAL PRODUCTS DIVISION

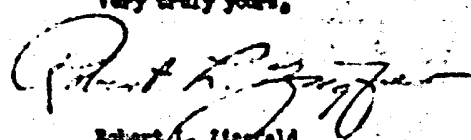
September 12, 1967

To Members of the Metallic Lead Products Division:

We are attaching a reprint from the Corrosion Forum of the "Chemical Engineering" magazine. This was sent to you as Chemical Bulletin No. 12. However, a correction to the last paragraph of the article has been added and was carried by "Chemical Engineering" in an addendum to the Corrosion Forum.

The attached reprint takes the place of Chemical Bulletin No. 12 and should be substituted for it.

Very truly yours,



Robert L. Zisgfeld
Acting Secretary

Wet and Dry Chlorine vs. Materials of Chemical Plant Construction

LEAD

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Lead is considered as a material for use in contact with chlorine gas up to 212 deg. F. with 10 percent moisture, and the amount of chlorine gas which can be used on lead, from concerns of chlorine, such as dry gas or water treatment, have no effect on lead. For handling hydrochloric acid, the use of lead is not generally recommended, but at normal temperatures there is only slight corrosion in concentrations up to 10 percent, and at 212 deg. F. concentrations as high as 20 percent have been used. Thus, when chlorine with some moisture present results in the formation of small amounts of dilute hydrochloric acid, there is a possibility of slight corrosion depending in amount upon other conditions.

Chlorine is a gas at room temperature and is inert to most metals when cool and dry, but the dry gas when heated attacks all metals to a certain extent. Hydrochloric and hypochlorous acids, formed in the presence of moisture, attack most metals except platinum, silver, lead, and certain special alloys. A protective layer of lead chloride forms over the lead and further corrosion is resisted while the film remains intact. This protective film permits lead's use with moist chlorine up to 210 deg. F. with only slight corrosion.

Lead pipe is used to introduce chlorine, gas or liquid, into a bath of caustic in a concrete tank in the manufacture of sodium hypochlorite. Pure lead tubing gives satisfactory length of service without undue expense. In general, at room temperatures, corrosion of lead by wet or dry chlorine, either liquid or gaseous, is sufficiently moderate to recommend

its use from an economic standpoint.

In nine years time, there has been approximately an eight per cent increase in the price of lead. From 40 cent per lb. to 48 cent per lb. of chlorine gas, the price of lead has increased 1.4 per cent. If chlorine gas is used in a system, it is essential that the chlorine is at approximately 125 deg. F. and is saturated with water vapor.

Relatively rapid corrosion takes place if lead contains a trace of chloride or is immersed in a relatively strong solution of a chloride and is subsequently exposed to the air. This is the result of a catalytic action in which the chloride acts as an intermediary with the formation of sodium carbonate and lead chloride which is subsequently attacked by the moisture and carbon dioxide in the air to form lead carbonate and sodium chloride.

Correction

In our original article for the symposium on chlorine (Chem. Eng., January 1947, page 211) the statement was made that lead is subject to rather rapid corrosion if exposed first to relatively strong chloride solution and then to the air. This is incorrect. The error dates back to a garbled 1912 abstract. While the reference is not directly pertinent to the subject, we feel an explanation is nevertheless in order.

It occasionally happens that lead objects in museums suffer in appearance from corrosion, with the formation of a white lead coating. C. Matignon in 1912 concluded that such corrosion might be attributed to sodium chloride in the air as a contributing factor. We believe other contributing factors to be: water vapor in the air, traces of organic acid vapors of the acetic acid type, oxygen and carbon dioxide.

H. M. CROSSLAND