

TABLE 1. GENERAL DATA OF COMBUSTIBLE ELEMENTS AND COMPOUNDS

SUBSTANCE	MOLECULAR SYMBOL	CHEMICAL REACTION OF COMBUSTION	IGNITION TEMPERATURE* Deg F	CALORIFIC VALUE				THEORETICAL OXYGEN AND AIR REQUIREMENTS			
				Btu per Lb		Btu per Cubic Foot	Higher	Lb per Lb		Cubic Ft per Cubic Ft	
				Higher	Lower			O ₂	Air	O ₂	Air
Carbon (to CO)	-	$2C + O_2 = 2CO$	-	4380	-	-	-	1.333	5.76	-	-
Carbon (to CO ₂)	-	$2C + 2O_2 = 2CO_2$	-	14540	-	-	-	2.667	11.52	-	-
Sulphur	S ₂	-	-	-	-	-	-	1.000	4.32	-	-
Sulphur (to SO ₂)	-	$S + O_2 = SO_2$	-	4050	-	-	-	-	-	-	-
Sulphur (to SO ₃)	-	$2S + 3O_2 = 2SO_3$	-	5940	-	-	-	-	-	-	-
Carbon monoxide	CO	$2CO + O_2 = 2CO_2$	1166-1319	4380	-	-	-	0.572	2.46	0.5	2.391
Methane	CH ₄	$CH_4 + 2O_2 = CO_2 + 2H_2O$	1202-1346	23850	21670	-	342	4.000	17.28	2.0	9.564
Acetylene	C ₂ H ₂	$2C_2H_2 + 5O_2 = 4CO_2 + 2H_2O$	763-824	21460	21020	-	1073	3.077	13.29	2.5	11.955
Ethylene	C ₂ H ₄	$C_2H_4 + 3O_2 = 2CO_2 + 2H_2O$	986-1123	21450	20420	-	1675	3.429	14.81	3.0	14.346
Ethane	C ₂ H ₆	$2C_2H_6 + 7O_2 = 4CO_2 + 6H_2O$	986-1123	22280	20500	-	1883	3.733	16.13	3.5	16.737
Hydrogen	H ₂	$2H_2 + O_2 = 2H_2O$	1063-1166	62000	52920	-	348	8.000	34.56	0.5	2.391
Hydrogen sulphide	H ₂ S	$2H_2S + 3O_2 = 2H_2O + 2SO_2$	599-608	-	-	-	-	1.412	6.10	1.5	7.173

*From International Critical Tables, 1927.

CHAPTER 8. COMBUSTION AND FUELS

ignition temperature before it will unite in chemical combination with oxygen to produce combustion. The ignition temperatures for several of the combustible constituents of fuels are presented in Table 1.

HEAT OF COMBUSTION

As previously stated, the process of combustion results in the evolution of heat. The heat generated by the complete combustion of a unit of fuel is constant for a given combination of combustible elements and compounds, and is known as the *heat of combustion*, *calorific value*, or *heating value* of the fuel.

The heat of combustion of the several fuel elements and compounds in their *pure* state is given in Table 1.

The reaction of the carbon in the fuel with oxygen may result in the formation of carbon monoxide or carbon dioxide. In burning to carbon monoxide, the carbon is not completely oxidized and, as shown by the data, the heat produced is considerably less than if it were completely oxidized. This fact is of greatest importance in considering the efficiency of combustion.

The calorific value of a fuel is determined by direct measurement of the heat evolved during combustion in a calorimeter. Although the ash and moisture content of coal from a given mine or locality may vary widely, the heating value of the coal, on a *moisture and ash free* basis, remains relatively constant. It is therefore possible to approximate the heating value of a shipment of coal *as received* if its moisture and ash content are determined, and if the heating value of similar coal on a moisture and ash free basis is known. This may be calculated by Equation 1.

$$\text{Heating value, as received,} = \frac{\text{Heating value, moisture and ash free} \times [100 - (\text{Moisture} + \text{Ash})]}{100} \quad (1)$$

where, moisture and ash are expressed in per cent.

The heating values for Illinois coals are published¹ and it is to be expected that values for other coals will be available in the future.

As practically all fuels contain hydrogen they produce a certain amount of water vapor as one of the products of combustion. The amount of water vapor produced increases as the hydrogen content of the fuel increases. When the heating value of a fuel is determined in a calorimeter the water vapor is condensed and the latent heat of vaporization that is given up during the condensation is reported as a portion of the heat value of the fuel. The heat value so determined is termed the *gross* or *higher* heat value and this is what is ordinarily meant when the heating value of a fuel is specified. In burning the fuel, however, the products of combustion are not cooled to the dew-point and the higher heating value cannot be obtained.

FLAME

The appearance of the flame or products of combustion may serve as an approximate measure of the temperatures developed in the combustion

¹State Geological Survey Bulletin, No. 62, Classification and Selection of Illinois Coals.