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528 Determination of T.N.T. (2,4,6-Trinitrotoluene) in Air. K. Kay. Canadian J. Res. 19, 86-9 (March, 1941).
War conditions have necessitated a reliable method for determination of T.N.T. in air. In the author's method the T.N.T. is absorbed in acetone, and 20% sodi hydroxide is added, producing the red sodium salt. The optimum amount of alkal was found to be 0.1 cc. per cc. of T.N.T. solution. The color developed to a maximum within 10 minutes and was stable up to 30 minutes. The intensity of re color was determined by a naked eye comparison with standard solutions. An accuracy of 10% over the range of 4-40 microgram per cc. of acetone could be obtained. With the aid of a photoelectric colorimeter and the proper color fil ter it would perhaps be possible to obtain much greater accuracy and sensitivit

529 Textile Bleaching Compounds. Industrial Data Sheet D-T1. Nat. Safety Council. Nat. Safety News, 43, 38, 40 (January, 1941).
The bleaching compounds discussed include sodium hypochlorite, hydrogen peroxid and sodium peroxide. Sodium hypochlorite may cause a dermatitis in the form of small blisters, particularly on the backs of the hands and between the fingers. Concentrated hydrogen peroxide, or even careless exposure to the 30% solution, may cause destruction or shedding of the finger nails, and bleaching of the hai Sodium peroxide is an irritant of the skin and mucous membranes. Burns from a hot solution are intensified by chemical action. Among the general safety sugg tions are: (1) use only by thoroughly trained workmen; (2) personal protective equipment; (3) flood-type emergency showers, and bubble type fountains for wash ing eyes; (4) use of closed systems when possible; (5) removal of supersensitiv persons from work with these materials; (6) chlorine gas masks for emergencies, and (7) care not to allow hypochlorite solutions and acid solutions to be pump into the sewer at the same time. First aid measures include liberal flushing c the affected skin with water, and neutralization of sodium hypochlorite and sodium peroxide with mild acids.

530 The Action of Methylene Blue upon Cyanide Poisoning in the Turtle. J.M.Johlin. J. Pharm. Exp. Therap. 71, 280-6 (March, 1941).
The heart glycogen of turtles receiving injections of methylene blue following those of sodium cyanide does not remain depleted as it does when they receive cyanide alone. The glucose and lactate contents of the blood, however, remain at high levels long after the administration of methylene blue or of sodium nitrite. The oxygen consumption of turtles poisoned by cyanide is very conside able but is greatly increased and brought within normal levels by methylene bl The amount of cyanide recoverable from the respiratory gases of turtles poison by cyanide is greatly decreased by methylene blue. These facts tend to support the theory that methylene blue converts hemoglobin into methemoglobin under th stabilizing influence of cyanides rather than that it takes the place of the respiratory enzymes of the cytochrome system of the respiratory mechanism.

531 A Case of Intoxication by Hydrocyanic Acid Following Disinfection by the Toxic Gas. M. Bornand and G. Bonifazi. Mitt. Lebensm. Hyg., 31, 207-10 (1940). (German).
Hydrocyanic acid has been used as a disinfectant for parasites in orchards and for rats and vermin in houses. It is more efficient than sulfur dioxide but more dangerous. Many deaths have resulted, and in Switzerland there are stric rules governing its use. The authors report a death which occurred on the 5th floor of a house when the first floor was being disinfected with hydrocyanic acid, due to an unnoticed shaft which connected all floors. The victim died within 20 hours, and at autopsy the greatest amount of hydrocyanic acid was found in the urine. Hydrocyanic acid should be allowed only in special cases and then under the supervision of an official disinfector.