

The history of development and validation testing of passive dosimeters is reviewed. Theoretical considerations, including possible limiting factors or interferences, are presented. Laboratory and field validation tests are critically reviewed and results are presented for comparative purposes. Evaluation of available data indicates that passive dosimetry, with some exceptions, is an acceptable method for monitoring gasses and vapors. Most importantly, passive systems appear to be as reliable as the now accepted active sampling systems.

Passive dosimetry — state of the art review

VERNON E. ROSE and JIMMY L. PERKINS

School of Public Health, University of Alabama in Birmingham, Birmingham, AL 35294

introduction

Recognition, evaluation and control are the cornerstones of the application of that mixture of science and art known as industrial hygiene. These three tasks, however, are no longer the eminent domain of the industrial hygienist. In the past decade, a proliferation of training in the recognition of workplace hazards has been made widely available to workers and management alike. At the other end of the spectrum has been the training of individuals highly specialized in the control of specific hazards, especially those involving noise and toxic air contaminants. These developments are welcomed because they contribute significantly to the ultimate goal of protecting the health of workers by providing safer and more healthful workplaces.

At the same time, professional industrial hygienists recognize that often the critical step in the process is not recognition of toxicity, but evaluation of hazard which leads to the subsequent development of the most effective means of control where warranted. This key step of evaluation is the unique domain of the industrial hygienist, often supplemented by other members of the occupational health and safety team. Where evaluation requires the determination of worker exposure to airborne toxic substances, the industrial hygienist has seen a revolution in the development of sophisticated techniques and equipment.

The "organ-grinder" impinger sampler is a relic, having been replaced by constant flow, eight-hour battery-operated pumps, light enough to be carried by the worker. The liquid bubbler and impinger have been replaced by the charcoal and chemical substrate sampling tube. The laboratory has come to the field in the form of the portable gas chromatograph and infrared monitor, albeit with a price rise directly proportional to the sophistication of the equipment. Even the once lowly, direct-reading detector tube has become legitimate with the establishment of government programs to certify accuracy and precision.

But while most evaluation techniques were reaching the point where the industrial hygiene staff required the addition of someone with a Ph.D. in electrical engineering, a new device has appeared which has the key of simplicity — the

personal passive dosimeter: personal, because it can be worn by the worker in close proximity to the breathing zone, passive, because there is no pump to move the air over a collector, which equates to fewer calibration and maintenance problems. Some quarrel with the term dosimeter, with purists preferring to call them collectors, monitors, or samplers. While many of the devices are collectors and require the application of subsequent analytical procedures, others provide for a more direct measurement of "exposure dose." Their basic appeal, however, is simplicity of use. Theoretically, elaborate calibration procedures are unnecessary, and all that is needed is a fairly reliable timepiece to measure exposure duration. There is some recognition that temperature and humidity may affect the observations, therefore, most manufacturers advise the user to report these environmental conditions to the analytical laboratory processing the dosimeter.

Rather than viewing passive dosimeters as another way to replace the industrial hygienist, industrial hygienists must recognize and appreciate the potential of the dosimeters in helping to achieve the hygienists' goals. That potential is significant in that personal dosimeters, if properly used, offer the opportunity to revolutionize the evaluation step. The parallels with detector tubes, as well as with noise and ionizing radiation dosimeters, are obvious. Indeed, the parallel with radiation dosimeters, especially film badges, is striking. The opportunity to significantly expand the measurement of worker exposure to many toxic materials can provide a quantum leap in our ability to provide safe and healthful workplaces. With any sampling device, however, there also must be the understanding that use of such devices is only one part of the evaluation step. The concepts of proper selection of workers at risk; the understanding of limitations, interferences and similar factors; and ultimately the proper interpretation of the results are still key ingredients in the evaluation step. The possibility of "false negative" decisions leading to erroneous assumptions of safety, or "false positive" conclusions leading to unwarranted expenditures of resources for controls, still exists regardless of the measurement device used.

With the rapid proliferation of passive dosimeters in the past several years, it is appropriate that industrial hygienists evaluate the "state-of-the-art" and, as professionals, become involved with the proper application of these monitoring devices.

theories of operation

In that passive dosimeters by definition do not use an air moving device to transport contaminated air to a collector, natural forces are relied upon to ensure that a representative amount of contaminant is "seen" by the detector. To date, one of two principles has been applied in the design of dosimeters. The first, and most widely used, is the principle of *diffusion* of contaminant molecules through a stagnant gas (air) layer. The second principle involves the absorption in and subsequent *permeation* of contaminant molecules through a membrane.

Diffusional monitors rely on the movement of contaminant molecules across a concentration gradient which for steady-state conditions, can be defined by Fick's First Law of Diffusion⁽¹⁾

$$W = -DA \frac{dc}{dx} \quad (1)$$

where, W = mass transfer rate, ng/sec.

D = diffusion coefficient, cm^2/sec .

A = cross sectional area of diffusion path, cm^2 , and

dc/dx = the instantaneous rate of change in concentration over diffusion path, $(\text{ng}/\text{cm}^3)/\text{cm}^{-1}$.

Considering the change in concentration ($C_1 - C_0$) over the total diffusion path length ($X_1 - X_0 = -L$), equation (1) becomes:

$$W = D \frac{A}{L} (C_1 - C_0) \quad (2)$$

where, L = length of the diffusion (static) path, cm.

C_1 = ambient concentration of contaminant, ng/cm^3 , and

C_0 = concentration of contaminant at collecting surface, ng/cm^3 .

If an effective collection medium is employed, the contaminant concentration at the surface of the collector (C_0) can be assumed to be zero, and multiplying both sides of equation (2) by time, yields:

$$M = D \frac{A}{L} (C_1) t \quad (3)$$

where: M = total mass transferred, ng, and

t = time that the badge is exposed to the contaminated air, sec.

It is also interesting to note that the units of the product of D and A , divided by L , are cm^3/sec , which are the same units associated with active air-moving devices such as personal sampling pumps.

Rearranging equation (3) as follows:

$$C_1 = \frac{ML}{DA t} \quad (4)$$

it becomes apparent that five factors affect the measurement of the ambient air concentration of a substance (C_1). Two of the factors (L and A) are physical parameters associated with the construction of the dosimeter, one (M) is provided by measuring the total mass of contaminant collected by the sampler, another is the duration (t) the sampler was exposed to the contaminated atmosphere, and the final factor (D) is an individual property of each vapor or gas. It also is known⁽¹⁾ that the diffusion coefficient is directly proportional to the absolute temperature (T) of the vapor, raised to three-halves power and inversely proportional to the atmospheric pressure (P).

$$D \propto \frac{T^{3/2}}{P} \quad (5)$$

Dosimeters that rely on the principle of *permeation* through a membrane are especially useful where the contaminant of concern is usually found mixed with other interfering vapors or gases or when a liquid collecting medium is employed. The goal then becomes to identify a membrane material that is highly permeable to the contaminant of interest and impermeable to most other components in the atmosphere, and/or the collecting media.

The determination of ambient concentrations of a contaminant using a permeation device can be determined from the formula:

$$C = \frac{w}{k t} \quad (6)$$

where: C = concentration of contaminant, ppm.

w = mass of contaminant collected, μg .

k = permeation constant, ppm-hours/ μg , and

t = exposure time, hours.

The permeation constant (k) is determined experimentally and is a function of the specific membrane material and contaminant of interest.⁽²⁾

sources of measurement error

The most obvious sources of error for both types of passive dosimeters are apparent from equations (4) and (6). Common to both badges are determinations of the mass of contaminant collected and the time of exposure of the dosimeter to the contaminated atmosphere. For the diffusional monitor, accurate knowledge of the physical parameters associated with badge construction (length and cross-sectional area) and the diffusion coefficient of the contaminant are important. There are at least nine prediction methods for calculating the diffusion coefficient, and in one study comparing observed and expected values for more than 100 compounds it was not uncommon to have less than 50 percent of the calculated results within ± 5 percent of the observed.⁽³⁾ Montalvo has described a procedure for limiting errors associated with computed diffusion coefficients.⁽⁴⁾ At

least one manufacturer, the 3M Company, makes available its procedures for determining sampling rates (DA-L) for its badges.⁽⁵⁾ Its approach has been to experimentally determine the sampling rate for five or six compounds in a chemical family to establish the relationship between the diffusion coefficient and the measured sampling rates. Sampling rates for other compounds are determined from the diffusion coefficients calculated by the Hirschfelder equation and the empirical relationships developed from the test compounds. The rationale for the selection of the Hirschfelder equation is not given, but in the study of the nine diffusion coefficient formulas, the author concluded that for higher molecular weight compounds the Hirschfelder, Bird and Spatz equations were in closest agreement with determined values.⁽³⁾ For the permeation monitor, accurate determination of the permeation coefficient for each monitor is necessary for obtaining accurate results. Factors influencing permeation include, thickness and uniformity of the membrane, affinity of the membrane for the analyte, swelling or shrinkage of the membrane, and possible etching by corrosive chemicals.

The problems associated with accurate determinations of the mass of the contaminant collected are similar to those involved with other collection devices such as charcoal or silica gel tubes, or to those in which the collection of the contaminant involves a chemical reaction with the collection medium. Using known amounts or concentrations of contaminants to determine collection and/or desorption efficiencies is as critical a step for passive dosimeters as it is for other methods of collection. Saturation of the sorbent as well as the subsequent accuracy of analytical techniques are also part of the total error associated with the measurement.

Another common concern in all types of environmental measurements is the potential for interferences, either positive or negative, from other contaminants in the sampled air. As the evaluation of passive dosimeters has matured, increased attention is being paid to possible interferences in multi-contaminant exposure situations, in both the laboratory and field. In evaluating such interferences it should be recognized that there are several potential sites for such interferences to appear, e.g., effects on adsorption or absorption efficiency of the sampling medium, chemical reactions of two or more contaminants prior to analysis, and the multitude of interferences associated with analysis of complex mixtures of gases and/or vapors. These problems also are found in the more classical sampling and analytical methods.

Accurate measurement of the time the sampling device is exposed, is essential to most industrial hygiene sampling procedures. For both short-term and full-shift exposure measurements, errors less than one percent, i.e. 9 seconds in 15 minutes and 4.8 minutes over 8 hours, are not unreasonable goals.

For the diffusion coefficient and possibly the permeation constant, it would appear that three factors have the greatest effect on variability. These factors are the two already identified, temperature and pressure, and, less readily apparent, the velocity of the air external to the badges.

Considering temperature and pressure, and referring to equation (5) it can be shown that a temperature rise from 5 to 35 °C would give a 16 percent increase in the diffusion coefficient, while a rise in barometric pressure from 710 to 810 mm Hg would cause a 14 percent decrease.⁽¹⁾ However, at the same time, the changes in temperature and pressure also are affecting the concentration (mass/volume; actually, density is the proper term but most authors use concentration) of the contaminant in that concentration is inversely proportional to the temperature and directly proportional to the pressure. As a result, the total mass (M) collected by the dosimeter is only slightly affected by temperature ($M \propto T^{1/2}$) and is independent of the pressure.⁽¹⁾ Consequently, while at ambient temperatures, the diffusion coefficient will increase about 0.5 percent per °C, and the total mass collected by the sampler will increase less than 0.2 percent per °C. Therefore, a temperature change from 25 to 30 °C, if uncorrected, will introduce a measurement error of less than one percent, while a change from 5 to 35 °C, if uncorrected, would introduce an error of about five percent.

The final source of error to consider is the velocity of the air external to the dosimeter; often this is referred to as face velocity. In an early assessment of face velocity effects, especially the lack thereof, Tompkins and Goldsmith point out that the important consideration is to contain all resistance to contaminant transport within the stagnant air layer inside the device.⁽¹⁾ As Jonas *et al.* subsequently noted, the face velocity directly affects the concentration gradient $C_1 - C_0$ in equation (2), and C_1 can no longer be assumed to be the ambient concentration when the air external to the badge is stagnant.⁽⁶⁾ With zero or low face velocities, the length (L) of the diffusion pathway is effectively extended, and there is a decrease in the measured ambient concentration. In Tompkins' and Goldsmith's work with the GASBADGE™, they determined experimentally that as long as face velocities were greater than 7.5 cm/sec (15 fpm) there was "no significant effect on dosimeter response;" however, experimental results supporting this conclusion were not presented.⁽¹⁾ High face velocities may also affect the concentration gradient. Commercially available diffusion devices rely on either a large ratio of diffusion path length to diffusion tube diameter or a wind screen to limit errors from this condition.

One of the most comprehensive tests to document sources of error has been conducted under contract for the National Institute for Occupational Safety and Health, and although concluded, it is not yet available as a public report.⁽⁷⁾ The study involved evaluation of the GASBADGE and 3M Organic Vapor Monitor™ (the DuPont badge not being available at the time the study was initiated) via challenge with several organic vapors. The factors investigated were precision, effects of storage, maximum and minimum levels of quantification, face velocity effects, effects of temperature and humidity, off-gassing (related to storage), exposure to mixtures, problems associated with applicable analytical methods, and adsorption of the contaminant by the badge itself with subsequent leaching to the sensing surface. The possibility of adsorption by the badge body, thus giving

higher results if the contaminant is subsequently released to the active medium, is of special concern in using passive dosimeters to measure very low ambient concentrations such as might be found in air pollution studies. Such interest and concern are evidenced by research on the subject being sponsored by the U.S. Environmental Protection Agency (EPA).⁽⁸⁾ Because of the lower concentrations involved with air pollution studies as opposed to workplace environments, the EPA also is concerned with the background or post-manufacture contamination levels associated with the sensing medium. Initially, the focus concerns organics and activated charcoal.

In summary, although numerous factors may affect the final calculation of concentration, only face velocity and the determination of the diffusion coefficient are unique sources of error for passive collectors. Therefore, if face velocities are sufficient to prevent "starvation" (probably greater than 7.5 cm/sec) and if diffusion coefficients have been accurately calculated or experimentally determined, passive dosimeters should give results comparable to those obtained with traditional active sampling systems.

statistical considerations

In evaluating any new monitoring method, extensive laboratory and field testing is necessary. Interpretation of the results of these tests requires the application of appropriate statistical techniques. The use of statistical techniques which are meaningful and easily understood is important; consequently, a discussion of the techniques used to evaluate passive dosimeters is appropriate.

There are numerous statistical tests which can be applied to both field and laboratory validation data. The main difference between the two situations is the degree of certainty of the "true" concentration of the monitored environment. In the field, the true value is usually an estimate based on the results of a standard sampling and analytical method. In the laboratory, experimental "known" concentrations are evolved and are then used for comparison with the concentrations estimated from sampling and analytical methods.

What is often not stated is that a certain amount of error also exists in the determination of laboratory-evolved "known" concentrations. These errors are often difficult to estimate. The "known" concentration is often calculated by weighing a syringe before and after an injection period (mass balance) or simply by injecting or allowing to diffuse a measured volume. It is assumed that the aliquot delivered was vaporized or diffused into a test chamber of known size. Possible sources of error include adsorption to or leaks from the test chamber, absorption and adsorption to articles placed in the chamber, degradation of the analyte by air oxidation or hydrolysis at high relative humidities, and error in measuring the injected contaminant. A backup monitoring system may be used to ensure close proximity to the "known" concentration. For example, an infra-red (IR) analyzer or gas chromatograph may be used as a check on a "known" concentration.

In other instances an IR analyzer or a direct reading instrument may be the only method for determining "known"

concentrations. In this case the error in the instrument can be calculated or estimated. Charcoal tubes and critical orifices also have been used to measure "known" concentrations. If error in both the "known" concentration and the estimated concentration are considered, statistical tests used to validate the experimental method become considerably more complicated; hence, the error in measuring the "known" concentration is usually assumed to be small and unimportant.⁽⁹⁾ Methods described above for determining the "known" concentration vary in their accuracy, a fact which should be considered when evaluating validation data for any sampling and analytical method.

When one is validating a method in the laboratory, there are two main considerations: the variation of the samples or data points about their mean, and the deviation of the sample mean from the true mean or "known" concentration. The first consideration often is called precision and is probably the most important and reliable measure as it does not depend on the error in determining the "known" concentration. Precision is estimated by determining the coefficient of variation (CV) or relative standard deviation of the data set as follows:

$$CV = \frac{s}{\bar{X}} \times 100 \quad (7)$$

where: s = Standard deviation of sample data set, and
 $\bar{X}$ = Mean of sample data set.

Where samples are taken at several concentrations, it is necessary to determine a pooled coefficient of variation which involves determination of number of levels tested.⁽¹⁰⁾ It should be noted that determining the number of concentrations (or more appropriately, the number of statistical levels) is not always straightforward. For example, if 10 samples are taken at each of three concentrations, and if within each concentration five of the samples are collected over four hours while the other five are collected over eight hours, are there three levels or six levels? The important point to consider is whether the differentiation of a level is based upon an anticipated difference in the sample mean. Certainly if the investigator designs the experiment with different time levels, there is an anticipated effect of time on the mean. Unfortunately, when experiments are so designed, statistical analyses at the various levels usually are not performed.

The second statistical consideration, the difference between the sample mean and the "known" value, is sometimes called accuracy, but the term bias is more appropriate. It is defined as:

$$b = \frac{\bar{X} - X_0}{X_0} \times 100 \quad (8)$$

where: $\bar{X}$ = mean of sample data set, and
 X_0 = "known" value at level tested.

If more than one level is sampled, it is necessary to determine the pooled bias of the data set.⁽¹¹⁾

The bias for a given set of data can sometimes be corrected. If the average bias (either the average of several samples at one level or the pooled bias for several levels) is large, one should note if the components of the bias value (either the individual samples or the levels) are consistently negative or positive. If the bias is large and varies consistently in one direction, the precision nevertheless may be quite small. In this case a physical or chemical variable may be consistently affecting the method (a systematic as opposed to random error), causing the experimental values to constantly fall short or long of the "known" concentration. This form of bias should be corrected.

In addition to these statistical tests, others have been used to assess the validity of passive monitoring systems. Overall system accuracy⁽¹¹⁾ has been defined as $(2 \times CV) + \text{absolute bias}$, expressed as a percent. Others⁽¹²⁾ have used the percentage of the "known" concentration accounted for by the sample mean $\pm$ two standard deviations as well as the term systematic error^(13,14) which is equivalent to overall system accuracy. Relative standard deviation has also been used, and is defined as the equivalent of CV.⁽¹⁵⁾ Additionally, some authors report only raw data while others report means without standard deviations or sample sizes, and various other combinations. While all of these statistical determinations have utility, it seems important for comparative purposes to consistently use those determinations which give the most information in the simplest form. Certainly, bias and precision meet these criteria.

Discussion of one other point seems necessary. NIOSH⁽¹⁰⁾ has proposed as a guideline for their own internal purposes that sampling and analytical methods meet a minimum requirement of ± 25 percent accuracy; that is, the absolute total error of the method should be less than 25 percent in at least 95 percent of the sample population (assuming a normal distribution). NIOSH derives the maximum precision value for an unbiased method given the accuracy criteria stated. This value (12.8 percent) is the maximum precision value acceptable for an unbiased method. Although the 25 percent accuracy criterion has been criticized by some authors,⁽¹⁶⁾ it was adopted for NIOSH's own internal use and is not meant as public policy. However, OSHA adopted the same criterion for the benzene standard, without a complete derivation or explanation. Consequently, this criterion has been criticized and alternatives have been proposed.⁽¹⁷⁾ A second important point is that bias is also considered in the ± 25 percent criterion according to a somewhat complex statistical relationship,⁽¹⁰⁾ but the overall system accuracy as defined earlier⁽¹¹⁾ is a fair approximation if the method has a true bias, i.e., its mean is statistically different from the "known concentration." A final point is that as the number of samples at a level increases, the standard deviation and CV should decrease. These considerations are important when evaluating passive monitor validation data, especially in those cases where manufacturers have stated that they have met the 25 percent accuracy criterion.

For field comparisons of conventional and passive monitors, different statistical tests are necessary. If passive monitor values, for example, are plotted against charcoal tube

results (Y vs. X) and more than one concentration is sampled, one would expect an increase in X to cause an increase in Y. If the increase is linear, and the sample values lie on the regression line, the correlation coefficient (r) would have a value of one. If a change in X brings about an equal change in Y, then the slope would also have a value of one. If the individual values for the two devices are indeed equivalent, the line should intercept the origin. Each of these relationships is expected within reason. The difference of the slope from one, the correlation coefficient from one, and the intercept from zero can and should be tested.

In order to perform the regression analysis described above we must assume that X (active sampling data) is not subject to error. Of course, we know and can calculate under laboratory conditions the error of active sampling systems. There are at least three reasons why this error is often overlooked. First, it is assumed that consideration of the error in X would only cause small differences in regression analysis results. Second, in addition to laboratory demonstrated error, a range of errors introduced by varying environmental conditions must be considered. While difficult to assess, these errors may have a profound effect on X, and indeed the error in the X measurement may be as great or greater than that for Y. Finally, if the error in X is to be considered the statistical tests are complex. Such tests have been discussed for biological problems,⁽¹⁸⁾ however, the theories apparently have not been applied to sampling and analytical methods even though their appropriateness has been recognized.⁽¹⁹⁾

applications

From a historical vantage, one of the earliest reports of a "passive" monitor for evaluation of airborne contaminants was patented by Gordon and Lowe in 1927.⁽²⁰⁾ Their gas detector for carbon monoxide involved an "easily frangible vessel containing a solution of salts including palladium chloride, and a covering for said vessel of a light colored absorbent material." The principle involved breaking the vessel (a small ampule) and noting the subsequent color change of the solution as it reacted with the carbon monoxide on the light colored absorbent material. This type of semiquantitative device was certainly a forerunner of those that are available today, though it was undoubtedly affected by air velocity as a stagnant air layer was not employed.

An extension of Gordon and Lowe's concepts in the late 1960's provided the basis for Plantz *et al.* to develop a personal dosimeter for measuring hydrazine, unsymmetrical dimethylhydrazine and monomethylhydrazine.⁽²¹⁾ The reaction of these compounds with a "colorimetric substance" (bindone) produced a purple color, the intensity of which was dependent on both concentration and duration of exposure. Color standards then were used to estimate the concentration as a function of the time the badge was exposed to the contaminated air; consequently, the method was only semiquantitative. In considering sources of error the authors noted that the purple color would also be produced by all volatile bases that were tested, including ammonia, aliphatic amines, aniline and cigarette smoke.

Of interest in this review, however, are *quantitative* devices based on the principle of either gas or vapor diffusion or permeation through a stagnant air layer. The first such device to be reported in the literature was described by Palmes and Gunnison in 1973.⁽²²⁾ Their device employed the principle of gas diffusion to determine airborne concentrations of sulfur dioxide and will receive further consideration subsequently. To gain the best overview of the various applications of these concepts, it is probably best to proceed by considering first the inorganic and then organic gases and vapors.

inorganic gases and vapors

ammonia

In 1978, Mazur *et al.* described the use of the Abcor GASBADGE to sample employee exposure to ammonia (this device currently is not marketed).⁽²³⁾ The investigators replaced the charcoal pad normally found in the GASBADGE with an acid impregnated absorption pad. Of three acids tested, phosphoric was most successful in providing the best approximation of theoretical concentrations. They determined, however, that volatile amines, specifically cyclohexylamine, could produce high readings, as high as 185 percent of the synthetic atmosphere. This led them to replace the glass fiber draft shield on the front of the GASBADGE with a "charcoal impregnated glass fiber filter which had been pretreated with alcoholic KOH containing 0.1 percent surfactant." The charcoal served to adsorb amines as they diffused into the dosimeter, while the KOH (aided by the wetting agent) eliminated irreversible ammonia adsorption by the charcoal, which would have caused underestimation of the ambient concentration. Additional laboratory experiments demonstrated that storage time of up to 47 days, prior to analysis, did not appear to adversely affect the results.

More recently, DuPont has developed a commercially available system for the measurement of several airborne contaminants including ammonia. In 1981, Kring *et al.* described DuPont's PRO-TEK™ system for ammonia, nitrogen dioxide and sulfur dioxide sampling analysis using a colorimetric readout instrument.⁽²⁴⁾ The ammonia badge relies on molecular diffusion of ammonia and subsequent chemical reaction with a solution of 0.3N boric acid and 0.03N sodium potassium tartarate (*sic*). After exposure, the reagent pack is removed from the badge holder and analysis is initiated by pressing reagent "blisters" which are adjacent to the absorbing solution. This action causes the release of a modified Nessler's reagent and the subsequent development of a colored solution. For ammonia, maximum color intensity is developed at 425 nanometers. The absorbance of the sample is then compared against a standard curve based on Beer's Law. After determination of the precision of the analytical method and verification of the linear range of the color chemistry, laboratory testing was conducted to establish the operational range as well as precision and accuracy of the overall method (see Table I). The minimum and maximum limits of the sampling range were found to be 50 and 500 ppm-hours, respectively. For an eight-hour time weighted average, these values correspond to one-fourth and two and

one-half times the current ACGIH Threshold Limit Value of 25 parts per million.⁽²⁵⁾ In considering sources of error, environmental effects including temperature (10 to 40 °C), relative humidity (10 to 80 percent), pressure (730 to 790 mm Hg), and face velocity (2.5 to 125 cm/sec) were included. Of the environmental factors evaluated, temperature and the concentration of the contaminant were identified as being responsible for 98 percent of the data variation. For ammonia, a temperature correction factor of 0.6 percent per degree centigrade was suggested. Also investigated was the storage stability of both unexposed and exposed badges. Results indicated that refrigerated storage is necessary to extend the shelf life of unexposed badges. Once the badge is exposed to ammonia and before the reagents are mixed, the badges can be stored for one (room temperature) to three (refrigerated) weeks without losing any absorbed contaminant. Once the reagents are mixed and color formation is started, the badge should be read within 90 minutes. Additional testing results conducted by DuPont are shown in Table I.⁽²⁶⁾

carbon monoxide

Shor and Anders, of the 3M Company, have described the 3M "direct-read diffusional monitor" for evaluation of exposures to carbon monoxide.⁽²⁷⁾ The principle involves the reaction of the carbon monoxide and an unreported reagent(s) to give a visible color change from pink to tan. Theoretically, "if any pink color is observable at the end of the exposure period, then the exposure was less than the one time-weighted-average of 400 ppm-hours." They also state that noting the time to the endpoint allows for calculation of the average concentration of carbon monoxide during the exposure period. The authors present summary results of laboratory evaluations using an infrared radiation device to establish "known" concentrations (see Table I).

chlorine

Hardy *et al.* have described a personal chlorine monitor (REAL, Inc.) which employs the principle of permeation of the gaseous contaminant through a silicone membrane and into 10 mL of a fluorescein-bromide solution.⁽²⁸⁾ Colorimetric techniques then can be applied to determine the chlorine concentration. The authors report a detection limit of 0.013 ppm chlorine for an eight-hour exposure, with a "working range" of 0.1 to 2.0 ppm. They also suggest that for shorter time periods, concentrations of up to five ppm can be determined. Other observations included effects of temperature, humidity, absorbent concentration, pH and response time, with all laboratory results presented graphically.

Moleculon Research Corporation has recently introduced a chlorine monitoring device which relies on plastic film impregnated with liquid reagents.⁽²⁹⁾ Exposure of the badge to chlorine gas gives a visible, "blue-purple," color change. Optical transmission measurements, and comparison with a standard curve, can then provide quantitative exposures in ppm-hours. The manufacturer's summary results indicating effects of temperature, humidity, wind velocity and concentration are reported as presenting an error at the 95 percent confidence level, which is "less than ± 15 percent."

TABLE I
Inorganic Gases and Vapors Laboratory Results

Chemical	Dosimeter ^a	Bias ^b	Precision ^c	Range ^d (ppm)	Reference	Notes
Ammonia	DP	0.5	7.4	20-50	24	
	DP	0.5	6.9	20-48	26	
	GB	-3.2	9.3	6-62	23	FG
Nitrogen	GB		21.7	4-11	1	H1
Dioxide	DP	-3.2	7.5	4-11	38	
	DP	-0.9	8.7	4-11	24	
	MDA	-4.9	4.1	6-9	36	FG
Sulfur	GB		13.8	4.6-5.3	1	H1
Dioxide	DP	-1	5.8	4-11	38	
	DP	-0.8	7.5	4-11	24	
Hydrogen Sulfide	GB		15.3	1.6-2.2	1	H1
Mercury	3M	0.5	9	0.03-0.3mg/m ³	32	F
	3M	17	2	0.05-0.2mg/m ³	33	FG
CO	3M	0.3	4.7	50-1830	27	G

^aDP=DuPont Pro-Tek Colorimetric System Badges, 3M=3M Company Monitor, MDA=MDA Scientific, GB=Abcor GASBADGE

^bSee text, equation (8)

^cSee text, equation (7)

^dSome values are rounded to nearest whole number

^eBias consistently negative

^fResults derived from Table III of McCammon *et al.*⁽¹²⁾

^gResults calculated from data provided in reference

^hThis product is not currently marketed

ⁱBias could not be calculated from data given

hydrogen sulfide

In 1977, Tompkins and Goldsmith described the development of the GASBADGE personal sampler.⁽¹¹⁾ Although later work with this device focused on the collection of organics on an activated charcoal substrate, initial studies researched sampling of both organic and inorganic compounds. (The GASBADGE for organic vapors is now marketed by National Mine Safety Company, and those for inorganic vapors are not currently marketed.) Applications involving inorganic gases relied on collective elements of an "appropriate substrate impregnated with a chemical medium specific for the contaminant of interest." Based on 60 observations, the authors' statistical summary is presented in Table I. Eighty percent of the measurements were within ± 25 percent of the "true" value. Challenge concentrations were established in a "well-mixed" environmental test chamber and were measured with an "independent wet-chemistry sampling train."

Hardy and associates have described a permeation device (REAL, Inc.) which relies on the permeation of H₂S through a dimethylsilicone membrane and subsequent reaction with a solution of 0.2N sodium hydroxide and EDTA.⁽³⁰⁾ The colored product (methylene blue) is measured spectrophotometrically and compared with a calibration curve to determine ppm-hours. A knowledge of exposure time then allows for the determination of average exposure over the measurement period. The authors note that a critical step in

the development of such a device is the experimental determination of the permeation constant (see Equation 6), which involves calibration of each monitor by exposure to known concentrations of the contaminant. The results of this laboratory research demonstrated a detection limit of 0.01 ppm for an eight-hour exposure, with a working range of 0.1 to 20 ppm and a linear response up to 200 ppm. The working range corresponds to one one-hundredth to two times the current eight-hour TLV of 10 ppm.⁽²⁵⁾ Evaluations of environmental effects indicated that neither temperature, over the range of -3 to 39 °C, nor humidity, from 0 to 99 percent relative, caused any significant variations in response of the device. Further research demonstrated a good response to high concentrations in less than one minute, adequate sample stability up to 10 days if EDTA is used in the absorbing solution, and negative and positive interferences, respectively, from chlorine gas and nitrogen dioxide. Precision and bias were not reported.

Another approach for the determination of gaseous hydrogen sulfide has been reported by Gracdel and Franey.⁽³¹⁾ Their research involved using a semiquantitative method without a stagnant air layer and relying on the discoloration of lead-stabilized polyvinyl chloride (PVC). The technique involves the diffusion of gas in a polymer and, at high H₂S levels, the detection rather than the measurement of toxic levels of H₂S. Screening applications for low level exposures also are discussed.

mercury

In 1977, McCammon and Woodfin of NIOSH reported the results of a laboratory evaluation of 3M's mercury vapor monitor.⁽³²⁾ The monitor's operating principle involves molecular diffusion and deposition of mercury vapor on a gold substrate. The resulting change in electrical conductivity across the gold foil is related to the amount of mercury absorbed by the foil. Three other sampling methods, all of which involved the active movement of air, also were investigated and include the LASL tandem sampling tube, the hopcalite tube, and the iodine impregnated charcoal tube. For the passive monitor, precision and accuracy, the effects of face velocity and temperature, and potential interferences were investigated. Concentrations of mercury vapor in an exposure chamber were monitored with an ultraviolet mercury vapor meter, which in turn was calibrated by measurements using the LASL method. To determine precision of the monitors, 12 devices were exposed to a test atmosphere. Results from three measurements of the test atmosphere using the LASL method gave an "expected" concentration of 0.056 milligrams of mercury per cubic meter of air (mg m^{-3}), with a standard deviation (SD) of 0.002 mg m^{-3} and a coefficient of variation (CV) of 0.037. Precision and bias calculated for the data given in Table III of McCammon *et al.*⁽³²⁾ is summarized in Table I. A least squares regression analysis of the combined precision and accuracy results for the passive dosimeters versus the "known" concentrations gave a Y intercept of -0.004 mg m^{-3} and a slope of 1.003. Tests for the effects of face velocities from 25 to 125 cm sec (50 to 250 fpm) did not appear to have any adverse effect on performance, while variation of temperature experimentally confirmed its effect on both the diffusion constant and the concentration of the contaminant. The potential effects of changes in relative humidity were not discussed by the authors. The other area of investigation in this study involved potential interferences of chlorine, sulfur dioxide, and hydrogen sulfide. Concomitant and sequential exposure to mercury and chlorine, the latter at high levels (5.8 ppm), produced a negative bias, but a similar effect was observed with the UV meter and the LASL system. The authors suggested that the mercury and chlorine may have been reacting to form mercuric chloride which was not being measured by any of the systems. The interference effects of sulfur dioxide and hydrogen sulfide, although less than chlorine, also were confirmed.

In 1980, McCammon and his NIOSH and OSHA co-workers conducted further tests of the four previously described mercury vapor sampling and analytical methods.⁽³³⁾ The design of this laboratory experiment involved inter-comparison of four methods, and the findings suggested that the variability of the iodine charcoal tube method is significantly different from that of the other three methods. It was observed that the other three methods, including the 3M passive dosimeter, exhibited good precision over the concentration range of 0.05 to 0.2 mg m^{-3} as shown in Table I.

Recently SKC, Inc., introduced a gas monitoring badge (produced by GMD, Inc.) which uses the principle of molecular diffusion (without a stagnant air layer) to collect mer-

cury vapor.⁽³⁴⁾ The sampling media is referred to as "Hydrar Sorbent," and "extensive" but unpublished field and laboratory testing are cited to show that "chlorine, moisture, etc., do not interfere" with measurements. HYDRAR is developed from a "manganese dioxide catalyst material similar to 'Hopcalite.'" Quantitative determination is made by chemical desorption of the mercury and analysis with atomic absorption.

nitrogen dioxide

In 1976, Palmes *et al.* reported the results of their evaluation of a personal sampler for nitrogen dioxide (NO_2).⁽³⁵⁾ This work was an extension of their earlier pioneering efforts in developing a personal sampler, employing the principle of gas diffusion, for sulfur dioxide.⁽²²⁾ In their design, the sampling device was a 1.3 cm (0.5 inches) acrylic tube, 7.1 cm (2.8 inches) long. At the "closed" end of the diffusion path (tube) were placed three stainless steel grids coated with triethanolamine (TEA). TEA was selected because, 1) it captures NO_2 efficiently, 2) it provides a stable sampling surface, and 3) it yields a chemical complex with NO_2 that is very stable over time. Subsequent analysis yielded a colored complex whose absorption was measured at 540 nanometers. Results were then compared against a standard curve which obeyed Beer's Law. The experimental evaluation of the NO_2 sampler involved chamber measurements compared with "known" values determined by the volume of NO_2 introduced to the chamber or the weight loss of NO_2 from a permeation tube. Although neither individual nor summary results were presented, graphical comparison of the data indicated a close agreement between the passive sampler results and theoretical concentrations. The effects of wind velocity and direction as well as stability over time also were considered. In determining wind effects, the uptake of water vapor, rather than NO_2 , was measured. The results indicated that there was an increase in average uptake with increased velocity and that the 45 degree incident angle gave the highest uptake (135 percent at 258 cm sec). Across all angles of exposure (0 to 180 degrees) the average uptake increased from 2 to 14 percent as the wind velocity increased from 50 to 258 cm sec (100 to 516 fpm). Stability studies indicated that the badges could be used for months both after preparation and before exposure as well as after exposure and before analysis.

In the previously described studies with the GASBADGE, Tompkins and Goldsmith also monitored for nitrogen dioxide.⁽⁴¹⁾ Their summary results for 82 observations showed accuracies and precisions as summarized in Table I. Eighty-five percent of the observations were within ± 25 percent of the "true" value.

A commercial model of the Palmes passive sampler has been marketed by MDA Scientific, Inc., and additional laboratory testing demonstrated a linear collection efficiency for any given dose, i.e., concentration $\times$ time (see Table I).⁽³⁶⁾ The sampler exhibited a consistently negative bias as compared to concentration determinations using a continuous monitor and the NIOSH wet chemical method.

As noted earlier, the DuPont PRO-TEK system includes a badge for nitrogen dioxide. Laboratory test results of this device are shown in Table I.⁽³⁷⁾

sulfur dioxide

As noted previously, pioneering work in 1973 on the design of a sampling device which relies solely on diffusion of gaseous contaminants through a stagnant air layer is attributed to Palmes and Gunnison.⁽²²⁾ Their initial studies involved experimental work on different tube lengths. The collecting medium was a complex of mercuric chloride and the final analysis involved colorimetric determination. The studies demonstrated that, except for very short diffusion paths (tube lengths), the diffusion monitor satisfactorily duplicated the results obtained by both wet chemical and conductimetric measurements. Although this study did not go into the ramifications of environmental effects and interferences, it should be recognized as an important step in developing a new industrial hygiene technology.

The last of the three inorganic gases looked at by Tompkins and Goldsmith in their studies of the GASBADGE

was sulfur dioxide.⁽¹⁾ Summary data for 23 observations on this gas are included in Table I. One hundred percent of the observations were within ± 25 percent of the "true" values.

The results from DuPont's tests on their sulfur dioxide badge also are shown in Table I.⁽³⁸⁾

organic gases and vapors

Most passive dosimeters designed to sample for organic gases and vapors use activated charcoal as the adsorbing medium. As we know from its extensive use in active systems, activated charcoal has an affinity for a wide range of organic compounds. Consequently, the discussion on the applications of passive dosimeters for the measurement of organic gases and vapors first will focus on devices using activated charcoal and then will turn to specific organics which rely on other collecting media.

TABLE II
Organic Gases and Vapors Laboratory Results

Chemical	Dosimeter ^A	Bias ^B	Precision ^C	Range ^D (ppm)	Reference	Notes
Carbon Tetrachloride	DPA	0.4	4.4	3-18	10	E
Toluene	DPA	0.3	5.25	57-228	45	FJ
	NMS	-1.7	1.7	12-47	13	
Formaldehyde	DP	1.5	6.3	0.2-4.2	55	
Benzene	DPB	3.3	4.7	3-24	46	FJ EGJ
	NMS	-4.1	1.7	0.8-5.4	13	
	NMS	1.8	17	13-13.5	1	
Ethylene Oxide	3M	-1.4	3.2	300	53	J
Halothane	3M	3	7.4	0.5-20	44	J
Enflurane	3M	-2.8	4.8	0.5-20	44	J
Acrylonitrile	NMS	0.03	8.7	0.7-19	39	
Hexane	NMS	-1.2	2	10-37	13	FJ
Vinyl Chloride	R	-0.2	3.7	1.5-14	2	HJ
Methyl Chloroform	3M	-5.9	4.7	160-840	48	IJ
Trichloroethylene	DPA	-10.6	4.5	160-840	48	EIJ
	NMS	2.4	2	15-65	12	FJ
	3M	-3.9	7.6	20-200	48	IJ
	DPA	-6.9	7.7	20-200	48	EIJ
	NMS	-0.5	1.9	15-67	12	FJ

^A3M = 3M Company Organic Vapor Monitor, DPA and DPB = DuPont PRO-TEK G-AA and G-BB Organic Vapor Badges, NMS = National Mine Safety GASBADGE, R = Real, Inc. MINIMONITOR, DP = DuPont PRO-TEK System Colorimetric Badges

^BSee text, equation (8)

^CSee text, equation (7)

^DSome values are rounded to nearest whole number

^ESmall sample size

^FKnowns were calculated using charcoal tubes with critical orifices. This could affect the bias measure.

^GPreliminary results

^HPermeation dosimeter

^IBias consistently negative

^JResults calculated from data provided in reference

activated charcoal devices

As of March, 1982, there were four manufacturers of passive dosimeters which rely on diffusion and subsequent adsorption on to activated charcoal. National Mine Service Company (GASBADGE), 3M Company (Organic Vapor Monitor), DuPont Company (PRO-TEK, G-AA and G-BB Organic Vapor Air Monitoring Badge), and the Mine Safety Appliance Company (Vaporgard Badge).

In 1977, Tompkins and Goldsmith described the first commercial passive dosimeter for monitoring organic vapors.⁽¹⁾ The GASBADGE relied on molecular diffusion of the vapor into the badge and subsequent adsorption onto activated charcoal. The authors developed the theoretical principles of the badge's operation, and discussed sensitivity to temperature and pressure, face velocity effects, and response time. Preliminary results showing the badge's response to benzene, ethyl acetate, methyl ethyl ketone, and styrene also were presented and were described as "very encouraging" (see Table II).

In the same year, Silverstein reported results of laboratory and field testing of the GASBADGE for acrylonitrile.⁽²⁾ The results of the exposure of 33 badges to known concentrations in the laboratory are presented in Table II. The results indicate the acceptability of the GASBADGE for measurements of acrylonitrile over the range of 0.75 to 19 ppm. The time of exposure in the laboratory was not given, however, field measurements did cover periods of up to seven hours. Although temperature and relative humidity ranges were reported, data analysis to determine the effects of these variables was not presented. Silverstein also looked at adsorption efficiencies and determined that the best results (4 percent) were obtained with four mL of two percent acetone in carbon disulfide.

In 1978, Bamberger *et al.* conducted a series of laboratory tests to evaluate the GASBADGE.⁽³⁾ Their approach involved the generation of known concentrations of solvents and subsequent evaluation with charcoal tubes (active sampling) and the passive dosimeter. To evaluate the applicability of the dosimeter over a wide range of compounds, the investigation included seven different organic compounds each representative of a different functional group. Included in this study were: benzene (aromatic), *n*-butanol (alcohol), *n*-butyl acetate (ester), isooctane (alkane), methyl chloroform (halogenated alkane), methyl isobutyl ketone (ketone), and trichloroethylene (halogenated alkene). The diffusion coefficient (*D*) used was that supplied by the badge manufacturer, except in the case of isooctane which was reported as having an unknown coefficient. Computations involving this compound relied on the coefficient for *n*-octane. A variety of experiments was conducted to look at a wide range of questions. Their findings corroborated dosimeter concerns similar to those of active systems using charcoal tubes, e.g., minimum and maximum loadings are important, post-sample contamination and loss can occur if the exposed absorbent is not adequately sealed, percent recovery for mixtures is consistent with percent recoveries for single compounds, and differences in charcoal lots can give different results. Other tests confirmed the need for some air

movement across the badge face and the lack of effect of temperature changes over a small range (11 °C). The results of the simultaneous sampling with the badges and the charcoal tubes indicated that the badges had a consistent negative bias. The authors suggested that, because the results were so reproducible, corrections for adsorption-desorption efficiencies less than 100 percent can be accomplished just as is done for charcoal tube data.

In 1979, Hirayama and Ikeda evaluated the application of the GASBADGE for monitoring exposures to mixed solvents.⁽⁴⁾ Their research involved different preparations of activated carbon "felt" in place of the supplied collection medium and exposure to mixtures of *n*-hexane, ethyl acetate and toluene. Summary (graphical) data indicated that the amounts of contaminant absorbed by the dosimeter were proportional to both the vapor concentrations and time of exposure.

Halliday and Anderson reported on the use of the GASBADGE in monitoring for halothane.⁽⁵⁾ Six observations indicated a range of measurements from minus nine to plus ten percent of the test atmospheres. Unfortunately, the authors did not report their procedure for determining the concentration of halothane in the test atmospheres.

In 1981, Evans and Horstman reported evaluations of desorption efficiencies of charcoal tubes and the GASBADGE for styrene.⁽⁶⁾ For liquid dosing they found the dosimeter to be similar to the tube, while for vapor dosing the badge was superior. The authors suggested that the differences in results may have been related to the use of coconut shell carbon in the tubes and petroleum derived carbon in the badge. They did not explain why this difference would affect one method of dosing and not the other.

In 1980, Anders and Mullins of the 3M Company presented results comparing the 3M passive monitor with charcoal tubes in sampling for mixtures of organic compounds.⁽⁷⁾ The laboratory tests included a binary mixture of toluene and methyl ethyl ketone; a tertiary mixture of benzene, toluene and xylene; and complex mixtures of unleaded and leaded gasoline containing various alcohols. Although the investigators cited "excellent" precision and accuracy for the diffusional monitor, sample sizes were small, and the complicated study design and lack of raw data preclude the determination of precision and bias statistics.

Mazur and his co-workers⁽⁸⁾ conducted side-by-side laboratory and field tests with charcoal tubes and 3M organic vapor monitors to measure concentrations of halothane (2-bromo-2-chloro-1,1,1-trifluoroethane) and enflurane (2-chloro-1,1,2-trifluoroethyl difluoromethyl ether). The results of the laboratory studies are presented in Table II and support the authors' conclusions that the dosimeters are a reliable method for the collection of enflurane and halothane.

In 1980, Lautenberger *et al.* described DuPont's passive monitor for organic vapors.⁽⁹⁾ Each charcoal strip in the PRO-TEK G-AA Organic Vapor Badge contains approximately 300 mg of coconut-based activated charcoal impregnated in an inert polymer. A dual sampling rate of approxi-

mately 50 or 100 ml/min is determined by the removal of one or both of the dosimeter's protective covers. One aspect of their research involved experimental determination of the diffusion coefficient of several gases and vapors. They reported that values calculated by Fugg⁽¹⁾ were within 10 percent of their experimentally determined diffusion coefficient values. Preliminary experimental results were used to discuss face velocity effects, range and sensitivity, maximum and minimum sampling times, vapor retention, storage stability, desorption efficiency, and overall badge efficiency. The overall accuracy determinations were limited to four observations at each of two concentrations of carbon tetrachloride (see Table II). However, the presentation of raw data, as well as an explanation of the statistical tests applied, is most useful. This same detail of information is also found in DuPont's validation reports for toluene and benzene (see Table II).^(17,18)

In the benzene report, DuPont also describes its PRO-TEK Ci-BB badge. This badge has a backup section of charcoal, which serves the same purpose as the second section in a charcoal tube, *i.e.*, to aid in determining if the sampler has been overloaded. The 3M Company also markets an Organic Vapor Monitor with a backup section.⁽¹⁷⁾

In studies of the measurement of waste anesthetic gases with passive dosimeters, Jonas *et al.* evaluated the GASBADGE, 3M Organic Vapor Monitor and DuPont Pro-Tek in measuring enflurane.⁽⁹⁾ Unfortunately, the badges were not identified in the presentation of the results although interpretation of the reported sampler geometry would indicate that A was the DuPont badge, B was the 3M badge, and C was the GASBADGE. The results of their laboratory studies indicated that badge B had the lowest coefficients of variation (CV was not calculated as described in this text) as compared to concentrations determined by infrared analysis. Badge A had a low CV (9 percent) at 5 ppm and a much higher value (CV=34 percent) at 20 ppm. The C badge had consistently high CV's ranging from 23 to 30 percent. It should also be noted that the charcoal tube CV's ranged from 11 to 27 percent, that desorption efficiencies for the badges ranged from 0.81 to 1.17, and that IR analyses of tank concentrations were constantly lower than expected. If badge B was the 3M device, the results of Jonas *et al.* support those reported by Mazur *et al.*⁽¹⁴⁾ Further testing of badges A and C seems necessary, however, to confirm their seemingly low precisions.

Mazur and his coworkers conducted additional tests comparing the 3M and DuPont badges against charcoal tubes.⁽¹⁸⁾ Methyl chloroform and trichloroethylene, two solvents widely used in vapor degreasing operations, were sampled. In the laboratory phase of the study, the badges and charcoal tubes were exposed to chamber concentrations over the range of 160 to 840 ppm of methyl chloroform and from 20 to 200 ppm of trichloroethylene. Exposure times varied from two to six hours for methyl chloroform and from four to six hours for trichloroethylene. The laboratory work indicated that the percent recoveries of the various doses (concentration $\times$ time) were in good agreement except for one exposure of the 3M badge which involved a five hour

exposure at 700 ppm. The authors noted that this exposure of 3500 ppm-hours exceeded the upper exposure limit provided by the manufacturer. The overall mean recovery value for each type of sampler was used to correct all subsequent field data. In addition to the recovery measurements, the laboratory phase of this study also involved determination of storage stability. The authors found no significant losses of methyl chloroform or trichloroethylene from either badge following storage of exposed badges for up to three weeks.

In 1978, West and Reiszner reported on the field tests of the MINIMONITOR™ (REAL, Inc.) permeation personal monitor for vinyl chloride.⁽²⁾ This monitor was a modified version of one previously described by Nelms *et al.*⁽¹⁹⁾ The collecting medium was activated charcoal, but rather than relying on molecular diffusion, the badge design involved a polymeric membrane and the permeation of vinyl chloride through the membrane and adsorption onto the charcoal. Initial laboratory calibration was used to determine the permeation constant of the device. Laboratory results indicated good accuracies as summarized in Table II.

During the same period that Tompkins and Goldsmith⁽¹⁾ were describing the GASBADGE, Bailey and Hollingdale-Smith of Great Britain were presenting their ideas for a personal passive sampler for organic gases and vapors.⁽¹⁰⁾ Their design involved the use of either one of two types of membrane and subsequent adsorption onto activated charcoal. They found two membranes to be satisfactory—one of thin silicone rubber which acted as a permeation barrier, and the second a porous polypropylene film which allowed for molecular diffusion of the gas and vapor. They conducted laboratory tests using carbon tetrachloride, styrene and dichlorodifluoromethane. Their test results do mix some terminology, *e.g.*, permeation rates for both the permeation device and the diffusional device, but did provide an early demonstration of the feasibility of such a device for monitoring certain organics. The device, the Porton Diffusion Sampler, seems to see its greatest use in Great Britain.

acrylonitrile

One of the newest applications of passive dosimetry involves the use of a porous polymer (Porapak® N) as the collecting surface with subsequent thermal desorption and gas chromatographic analysis. Benson and Boyce have described such a device and its utility in sampling for acrylonitrile.⁽²¹⁾ Laboratory testing for acrylonitrile involved comparison of the dosimeter values with concentrations measured on a gas chromatograph. Initial experimentation indicated that the dosimeter can be used for acrylonitrile concentrations in the range of 4 ppm, but at concentrations of 2 ppm a 40 percent error is reported.

aniline

In addition to activated charcoal, another widely used adsorbent medium is silica gel. To study the utility of this material, Campbell and Konzen constructed passive dosimeters from glass culture tubes (1.05 cm inside diameter) with 40-60 mesh silica gel as the collecting surface.⁽⁵²⁾ Laboratory testing involved exposure of the dosimeter to aniline, with exposure concentrations determined by gas chromatographic analysis of ethanol gas scrubbers. Three different

size (length) dosimeters were evaluated, with the best results obtained with the intermediate length tube ($L = 3.0$ cm; $A/L = 0.3$ cm). The authors present raw data and clearly described their statistical techniques.

ethylene oxide

Mullins and Anders have recently described the 3M diffusional monitor for sampling ethylene oxide in air.⁽⁵³⁾ In this badge the collecting surface is described as a "chemically impregnated charcoal surface, (where) a reaction occurs producing a stable compound with a vapor pressure substantially lower than the parent compound." The authors present statistically summarized data describing the linearity and capacity of the monitor, the recovery of absorbed ethylene oxide, environmental effects, sample stability, and the effects of potential interferences. Precision and bias are presented in Table II.

formaldehyde

Rodriguez *et al.* have described another 3M diffusional monitor for sampling formaldehyde.⁽⁵⁴⁾ In this diffusional monitor, the collecting surface is an "impregnated sorbent" which can then be desorbed *in situ* with water and the concentration of formaldehyde determined colorimetrically. Laboratory evaluation first involved determination of recovery coefficients, which at eight ppm-hours (19.5 micrograms) were found to be 1.00 ± 0.04 over six tests. The next step involved determination of the dosimeter's "sampling rate" (DA 1) by exposing the dosimeters to "known" concentrations of formaldehyde as generated by a permeation tube. The effect of relative humidity on the sampling rate also was investigated, and evaluation of the data did not indicate any statistically significant differences between the rates at 50 and 85 percent relative humidity. The study protocol then involved simultaneous exposures of impingers (modified chromotropic acid method) and dosimeters. The authors concluded that "the measured values by both methods lie within ± 25 percent of the expected response" and that "less variation is observed in the monitors than in the impingers." However, neither precision nor bias were reported. The authors also investigated effects of storage and determined that at elevated temperatures (38°C) losses up to 11 percent occurred after one week, however no significant loss was seen for samples stored at 23°C . The authors briefly discussed the potential for a negative interference from phenol and described the use of modified calibration curves to address this problem.

DuPont's PRO-TEK series of colorimetric Air Monitoring Badges, includes a badge for formaldehyde. The collection principle involves a chromotropic acid-sulfuric acid reaction. Laboratory evaluation (42 samples) of the device at seven exposure levels revealed results as shown in Table II.⁽⁵⁵⁾ Additional studies also were conducted on temperature and storage effects. The raw data and statistical analysis procedures are presented.

Kriesel⁽⁵⁶⁾ has described a new passive dosimeter for formaldehyde which is a modified version of the Palmes tube. At the present time experimental data concerning this device are not available.

phosgene

Matherne *et al.* have recently described the GMD, Inc. "passive dosimeter" which provides a semiquantitative measurement of phosgene exposure.⁽⁵⁷⁾ The badge involves direct contact between the contaminated air and a chemically impregnated tape and therefore does not rely on a stagnant air layer. The treated paper stain intensity is reported to be logarithmically proportional to the phosgene dose over a range of 2 to 100 ppm-minutes. For quantitative measurements the badges can be read colorimetrically.

other methods

Hill and Fraser have described the use of commercial detector tubes modified to act as passive dosimeters.⁽⁵⁸⁾ In their research, common length-of-stain detector tubes were modified by cutting off the conical end of the tube and removing some of the indicator column material. This leaves an orifice with a cross-sectional area equal to that of the inside of the tube and a path length determined by the distance from the end of the tube, to the beginning of the indicator material. One would expect, however, that as the sorbent material becomes exposed, *i.e.*, the length of stain increases, the diffusion path length will also increase, thereby changing the sampling rate. Their evaluation of these devices involved separate laboratory exposures to toluene, ethanol and isopropanol. The results of their work, although presented only in graphical summary, demonstrate the potential for the use of modified commercial detector tubes as passive dosimeters.

field validation

Relatively few studies have been published in which passive dosimeters have been compared side by side with charcoal tubes or other conventional sampling methods under actual field conditions. For inorganic compounds only two studies, involving nitrogen dioxide⁽⁵⁹⁾ and chlorine,⁽⁶⁰⁾ have been identified. For organic compounds, eight studies^(2,9,39,44,48,51,60,61) have involved field comparisons, with the number of compounds per study ranging from one to 22. In three of these studies, statistical analyses of data were not presented and cannot be performed because of small sample size or insufficient presentation of data.

Jones *et al.*⁽⁵⁹⁾ conducted a field evaluation for NO_2 in a salt mine, which contained diesel equipment as the NO_2 source. At each of 16 different fixed area locations, two Palmes dosimeters and two TEA tubes with pumps were used to sample the atmosphere. The active sampling (pump) method gave a coefficient of variation (CV) of 8.7 percent while for passive tubes the CV was 5.8 percent. Regression analysis (where the active system was the X variable) of their data gives a correlation coefficient of 0.69, a slope of 0.59, and an intercept of 2.05 ppm for data over the range of 3.7 to 5.5 ppm as sampled by the active method. This indicates that the dosimeter gave consistently higher readings which is reflected in the means for the two methods: 4.51 ppm for passive and 4.14 ppm for active sampling. The authors did not report wind velocities, but low velocities, as would be expected with area samples, should have caused passive

values to be low as compared to active values: this was not the case. On the other hand, high face velocities could have led to the observed positive (passive versus active) bias.

Hardy *et al.*⁽²⁸⁾ reported the raw data results from a field evaluation of a permeation chlorine monitor (REAL, Inc.). Thirteen comparisons were made in which the results from a battery operated pump and an impinger sampler were compared with the measurements from either two or three permeation samplers. To further evaluate their results, the investigators performed a regression analysis using their permeation sample means for each comparison as the dependent variable. The results are quite good with a correlation coefficient of 0.95, a regression slope of 0.85, and an intercept of 0.15 over a range of 0.03 to 1.1 ppm as detected by the impinger. It should be noted that with five impinger samples of less than 0.1 ppm, the corresponding permeation devices detected considerably higher concentrations (0.16 to 0.4 ppm).

Silverstein⁽³⁹⁾ reported field results for acrylonitrile monitoring using 18 paired samples of GASBADGE passive monitors and active systems (charcoal tubes and pumps) over a range of 0.8 to 3.8 ppm as determined by the active method. The differences in results using the active system as a reference ranged from -0.7 to 1.5 ppm. The difference in means, 2.18 for the passive versus 2.73 for the active system, was 25 percent. Further data were not presented.

West and Reiszner reported five sets of field results for vinyl chloride sampled with permeation dosimeters (REAL, Inc.) and charcoal tubes.⁽⁴²⁾ Further interpretation of their results is presented in Table III. In each case data for the active system are the λ values. Four of the correlation coefficient values are very near one, reflecting good correlation, however, the slopes show quite a large degree of variability (0.69 to 1.31), indicating that badge values may fall either well below or above charcoal tube values. In those cases where the slopes were less than 1.0, very high humidities (67 to 91 percent) were reported by the authors. This factor may have interfered with permeation, although the authors reported that humidity had no effect in laboratory validations, hence the variation in slope remains unexplained. The authors noted that the overall field results showed that the badges had a slight positive bias.

Hickey and Bishop exposed 78 pairs of side-by-side charcoal tubes and 3M Organic Vapor Monitors to complex mixtures of organic chemicals in tire manufacturing operations.⁽⁹⁾ Generally the sampling period ranged from three to

five hours, and most observations consisted of one monitor and the time weighted average concentration from two sequentially exposed charcoal tubes. Sixty-four of the sets were personal samples, while the remaining 14 were area samples. The samples were collected in two separate plants (30 sample pairs in one plant and 48 in the other). Of the 22 organics potentially available for analysis, 10 were detected over a sufficiently wide range of concentrations to allow for appropriate statistical analysis by linear regression. The results are interesting in that in the first plant, 9 of the 10 organics measured by the dosimeters showed higher vapor concentrations as compared to the charcoal tubes, while in the second plant only three substances had a regression slope greater than one. The combined data for both plants did not indicate that the passive system was consistently biased when compared to the active system. The authors did point out that generally the Y-intercepts (Y =passive dosimeter data) were slightly negative, a finding which may indicate a lack of sensitivity on the part of the dosimeters at low concentrations. For the remaining 12 compounds, paired t-tests revealed no significant difference between the charcoal tube and passive monitor means at the 95 percent confidence level. The use of t-tests to analyze such data has been questioned since the means of the two methods may be very similar but the components of paired values can be considerably different.⁽¹⁹⁾ This condition can only be revealed through regression analyses.

In 1980, Mazur *et al.*⁽⁴⁴⁾ reported limited field data for halothane and enflurane measurements using both 3M Organic Vapor Monitors (OVM) and an active system (charcoal tubes and pumps). For halothane three paired samples were reported. The mean concentration for the active system was 2.01 ppm while 1.9 ppm was reported for the OVM, a difference (relative to the active system) of five percent. Only one data pair was reported for enflurane: 0.49 ppm for the OVM and 0.52 ppm for the active system. Obviously, more data are needed to draw conclusions regarding a comparison of the two methods for these agents.

A second study by Mazur *et al.*⁽⁴⁸⁾ reported field comparisons of passive dosimeters and active systems (pumps and charcoal tubes) in sampling for trichloroethylene (TCE) and methylchloroform (MC). Both DuPont PRO-TEK and 3M Organic Vapor Monitors were used for the passive systems. Personal samples included exposure of one each of all three monitors. Area sample results involved three average values: one was the average of three charcoal tubes, and the other two, the average of two of each type of dosimeter. For MC, 11 personal and 7 area data points collected over time periods of 1 to 5 hours at 15 to 21 °C and 35 to 40 percent relative humidity were reported. For TCE, 22 personal and 7 area data points collected over periods of about 1 to 6 hours at 18 to 24 °C and 30 percent relative humidity were reported. A regression analysis in which the charcoal tubes were the independent variable was reported by the authors. In each of the following data sets the presented values involve TCE personal and stationary sampling followed by MC personal and stationary sampling. For the DuPont badge, regression

TABLE III
Regression Analysis of
Vinyl Chloride Field Data⁽⁴²⁾

N	Range (ppm)	r	Slope	Y-Intercept
7	0.02-1	0.99	1.19	0.03
8	0.08-1.8	0.99	0.69	0.05
12	0.02-6.9	1.0	1.31	0.01
39	0.05-1.8	0.82	0.81	0.11
24	1.48-16.7	0.96	1.08	0.43

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