Table 1. General Data of Combustible Elements and Compounds[a]

Substance	Molecular Symbol	Chemical Reaction of Combustion	Ignition Temperature F Deg	Calorific Value: Btu per Lb, Higher	Calorific Value: Btu per Lb, Lower	Calorific Value: Btu per Cu Ft, Higher	Theoretical Oxygen and Air Requirements: Lb per Lb, O_2	Lb per Lb, Air	Cu Ft per Cu Ft, O_2	Cu Ft per Cu Ft, Air
Carbon (to CO)	—	$2C + O_2 = 2CO$	—	3950	—	—	1.332	5.763	—	—
Carbon (to CO_2)	—	$2C + 2O_2 = 2CO_2$	—	14093	—	—	2.664	11.527	—	—
Sulphur (to SO_2)	—	$S + O_2 = SO_2$	—	3983	—	—	0.998	4.285	—	—
Sulphur (to SO_3)	—	$2S + 3O_2 = 2SO_3$	—	5940	—	—	1.497	6.428	—	—
Carbon Monoxide	CO	$2CO + O_2 = 2CO_2$	1166-1319	4347	—	321.8	0.571	2.471	0.5	2.382
Methane	CH_4	$CH_4 + 2O_2 = CO_2 + 2H_2O$	1260-1380	23879	21520	1013.2	3.990	17.265	2.0	9.528
Acetylene	C_2H_2	$2C_2H_2 + 5O_2 = 4CO_2 + 2H_2O$	763-824	21500	20776	1499	3.073	13.297	2.5	11.911
Ethylene	C_2H_4	$C_2H_4 + 3O_2 = 2CO_2 + 2H_2O$	986-1123	21644	20295	1613.8	3.422	14.807	3.0	14.293
Ethane	C_2H_6	$2C_2H_6 + 7O_2 = 4CO_2 + 6H_2O$	990-1120	22320	20432	1792	3.725	16.119	3.5	16.675
Hydrogen	H_2	$2H_2 + O_2 = 2H_2O$	1063-1166	60958[c]	51571[c]	325	7.937	34.344	0.5	2.382
Hydrogen Sulphide	H_2S	$2H_2S + 3O_2 = 2H_2O + 2SO_2$	599-608	7100	6545	647	1.409	6.097	1.5	7.146
Propane	C_3H_8	$C_3H_8 + 5O_2 = 3CO_2 + 4H_2O$	950-1080	21661	19944	2590	3.629	15.703	5.0	23.821
n-Butane	C_4H_{10}	$2C_4H_{10} + 13O_2 = 8CO_2 + 10H_2O$	890-1020	21308	19680	3370	3.579	15.487	6.5	30.967
Commercial Propane	—	—	—	21650	—	2500	—	—	—	—
Commercial Butane	—	—	—	21400	—	3200	—	—	—	—

[a]Values in table taken chiefly from page 51 of Fuel Flue Gases published by *American Gas Association*. [b]Gas measured at 60 F and 30 in. Hg.
[c]Value from *National Bureau of Standards*.

The heat of combustion of the several fuel elements and compounds in their *pure* state is given in Table 1.

The reaction of the carbon in the fuel with oxygen may result in the formation of carbon monoxide or carbon dioxide. In burning to carbon monoxide, the carbon is not completely oxidized and, as shown by the data, the heat produced is considerably less than if it were completely oxidized. This fact is of greatest importance in considering the efficiency of combustion.

The calorific value of a fuel is determined by direct measurement of the heat evolved during combustion in a calorimeter. Although the ash and moisture content of coal from a given mine or locality may vary widely, the heating value of the coal, on a *moisture and ash free* basis, remains relatively constant. It is therefore possible to approximate the heating value of a shipment of coal *as received* if its moisture and ash content are determined, and if the heating value of similar coal on a moisture and ash free basis is known. This may be calculated by Equation 1.

Heating value, *as received* =

$$\frac{\text{Heating value, } \textit{moisture and ash free} \times [100 - (\text{Moisture} + \text{Ash})]}{100} \quad (1)$$

where, moisture and ash are expressed in per cent.

Typical analyses of the coals of the United States are given in *U. S. Bureau of Mines Bulletin* 446.

As practically all fuels, including coal, oil, and gas, contain hydrogen, water vapor is one of the products of combustion. The amount of water vapor produced is proportional to the hydrogen content of the fuel.

When the heating value of a fuel is determined in a calorimeter the water vapor is condensed and the latent heat of vaporization is included in the heating value of the fuel. The heating value so determined is termed the *gross* or *higher* heating value and this is what is ordinarily meant when the heating value of a fuel is specified. In burning the fuel, however, the products of combustion are not cooled to the dew-point and the higher heating value cannot be obtained.

FLAME

The appearance of the flame or products of combustion may serve as an approximate measure of the temperatures developed in the combustion process. The luminosity of a flame is caused by the heating to incandescence of unconsumed particles of combustible matter in the gases, and the higher the temperature of these particles the whiter the flame. Table 2 gives some approximate flame temperature data.

Table 2. Approximate Flame Temperature Data

Appearance of Flame	Temperature F Deg
Red, visible in daylight	1000
Light red	1800
Orange-red	2000
Orange-yellow	2200
Yellow-white	2400
Bright white	2600