

ENVIRONMENTAL UPSET REPORTSUBJECT: LEAK AT OHC II CATALYST CATCH POTFROM: DON PEARSON

Please complete the following questionnaire and forward as soon as possible.
[This report must be received by the Environmental Control Department within five (5) calendar days of the incident.]

1. DATE OF INCIDENT: 2-25-83
2. SOURCE OF RELEASE: HOLE IN OHC CATALYST CATCH POT
3. TIME AND DURATION: 8 AM - 11 AM
4. WEATHER CONDITIONS (Circle One): Clear Partly Cloudy Cloudy Rain
5. WIND SPEED AND DIRECTION (if a gas release): 10 mph N
6. HAZARDOUS MATERIAL RELEASED: HCL, EDC, CHLORAL, TCE
7. ESTIMATED QUANTITY RELEASED: 385 lbs
8. CAUSE OF RELEASE: HOLE IN CATALYST CATCH POT

9. IMMEDIATE ACTION TAKEN: REDUCED REACTOR RATE TO MINIMUM
TO TRY TO PLUG LEAK. WHEN IT WAS DECIDED
LEAK COULD NOT BE STOPPED, REACTOR WAS SHUT DOWN.
10. ACTION PLANS TO PREVENT REOCCURRENCE: WILL TAKE A CLOSE LOOK
TO DETERMINE CAUSE OF FAILURE. CATCHPOT WILL
BE TEMPORARILY REPLACED WITH A STEAM-TRACED APE
SPOOL AS WAS DONE IN NOS. 3, 5 + 6 REACTORS.
11. NOTIFICATION MADE (Who, When): PLANT B II SHIFT ENGINEER

Environmental File # _____

SL 047767

CALCULATION

Loss through oval hole $\frac{1}{4}$ " wide by $\frac{1}{2}$ " long
 $\sim$.30" dia round hole

Process pressure was changing as reactor rates were reduced and reactor was shut down. Figure average pressure was 15 psig and material was lost over a period of 3 hrs. (8 AM to 11 AM)

Composition of reactor vent: (wt%)

O ₂	0.49
N ₂	0.70
CO	0.27
CO ₂	8.76
C ₂ H ₄	7.83
HCl	0.63
VC	0.01
EC	0.01
DCE	0.06
EDC	67.75
CCl ₄	0.09
CHLORAL	0.37
TCE	0.84
H ₂ O	12.19

Average MW = 52.5

To calculate loss use eq. from Crane p.3-5

$$q = .525 d_1^2 C \sqrt{\frac{\Delta P}{\rho}}$$

$$\Delta P = 15 \text{ psig} \quad \rho = \frac{52.5 \text{ lbs}}{\text{mole}} \times \frac{\text{mole}}{379 \text{ ft}^3} \times \frac{29.7}{14.7} \times \frac{520}{910^\circ}$$
$$= 0.160 \text{ lb/ft}^3$$

$$d_1 = \text{orifice diameter} = 0.3 \text{ in}$$

from pp A-20 Crane use maximum value
of C for $\beta = 0$ (to be conservative)
 $\Rightarrow C = 0.7$

$$q = .525 (.3)^2 0.7 \sqrt{\frac{15}{.160}} = 0.320 \frac{\text{ft}^3}{\text{sec}}$$

$$0.320 \frac{\text{ft}^3}{\text{sec}} \times \frac{3600 \text{ sec}}{\text{hr}} \times \frac{0.160 \text{ lb}}{\text{ft}^3} \times 3 \text{ hrs} =$$

552 lbs lost over 3 hrs

Hazardous materials lost:

HCL	=	3.5 LBS
VC	=	0.06 LBS
EDC	=	374 LBS
CHLORAL	=	2 LBS
TCE	=	5 LBS

Non-hazardous materials lost: N_2 ,
 C_2H_4 , H_2O make up the balance
of the 552 lbs.