Anniston, Alabama

March 2, 1948

Messrs.
R.L. Jenkins - C.B. Durgin
Paul Logue - Ph.Sls. St. Louis
F.A. Barker (2)
R.L. Wolf
A.M. Ellenburg (3)

Mr. P.G. Benignus
Ph. Sls.
St. Louis

We are greatly enjoying the numerous, interesting, and detailed Sales Development Reports which we have been receiving from you recently. I regret that we have been unable to keep up with all the resulting need for additional information and with the considerable research work which will be required to make fullest use of the available information which you have accumulated. In order to show you that we intend to make the fullest use of the data obtained, I thought it might be good to comment on some of your reports, even though we are not actually completing or in some cases, even starting the desired experimental work.

Edgar E. Hardy

EEH:rs

Attached

DSW 330891

STLCOPCB4077535

Advance Solvents and Chemical Corporation
245 Fifth Avenue
New York, New York

Mr. Gerry P. Mack, Director of Research

HB-40

The problem of light stability of HB-40 has been of much concern to us for a considerable time and a great deal of experimental work has been done. Concerning the variations in stability of HB-40 from lot to lot, I believe that we should institute a routine color stability test and we shall develop one if at all possible. We have known that unstable samples tend to develop objectionable odor.

We shall obtain a sample of nonyl naphthalene and check its properties. We might prepare a similar material by alkylating biphenyl or terphenyl.

We agree with Mr. Mack that HB-40 does not respond favorably to the action of commonly used stabilizers.

Concerning the three materials of improved stability which were tested by Advance Solvents, sample S.F. was a fractionated HB-40, sample S-4 was overhydrogenated and S-1 was underhydrogenated. The complaint that sample S.F. which looked most promising precipitated and also showed a lower boiling range than HB-40 might be overcome by additional research work.

We have at present a series of hydrogenations underway, aiming at the preparation of material equivalent or better than the HB-40 sample S.F. We shall inform you of the results of these tests as soon as they are completed.

Dr. Jenkins intends to ask Dayton whether they would be able to contribute to the solution of the problem of stabilizing HB-40 by possibly attacking it from a new and original viewpoint. Additional work in Anniston will await the results of the tests now in progress.

I believe that you have since determined that there is a steam coil in the HB-40 tank car so that this point has been clarified.

We agree with you that Aroclor 1254 should certainly be stable at 160°C. and not evolve HCl at this temperature.

DSW 330892

STLCOPCB4077536

Aerovox Corporation
Belleville Avenue
New Bedford, Massachusetts

Aroclor 1254

As requested in your memo of February 12, 1948, we are enclosing the detailed procedures for the determination of the dielectric constant and resistivity of liquid Aroclors as practiced in the Anniston Analytical Laboratory.

Concerning the fluorescence test, we believe that while they may not be much concerned whether we run this test or not, we should definitely obtain, if at all possible, the detailed test procedure as to the type of container and lamp used. We will then use the procedure so that we might not encounter a refusal of a shipment someday, based on this test.

DSW 330893

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